



UNIVERSITI PUTRA MALAYSIA

***SOLUTION-ENHANCED DISPERSION BY SUPERCRITICAL CARBON
DIOXIDE IN PRODUCTION OF *Andrographis paniculata* (Burm.f.) Wall.
Ex Nees POWDER***

LEE SIN YEE

FSTM 2019 6



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By

LEE SIN YEE

**Thesis Submitted to the School of Graduate Studies, Universiti Putra Malaysia, in
Fulfillment of the Requirements for the Degree of Doctor of Philosophy**

November 2018

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Abstract of thesis presented to the Senate of Universiti Putra Malaysia in fulfillment of the requirement for the degree of Doctor of Philosophy

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November 2018

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Andrographolide with poor aqueous solubility and poor dissolution rate is predisposed to low oral bioavailability. To overcome these constraints, strategies (reducing crystallinity and increasing surface area) were employed through Solution-Enhanced Dispersion by Supercritical Carbon Dioxide (SEDS). SEDS precipitation of *Andrographis paniculata* (Burm.f.) Wall. ex Nees extract from CO₂-Acetone and CO₂-Acetone:Ethanol (v/v) 1:1 systems at different pressure (100, 150 bar) and temperature (40, 50 °C) combination as well as aqueous solubility of andrographolide precipitated were first studied. Modification of its aqueous solubility was then conducted by manipulating precipitation using CO₂-Acetone:Ethanol (v/v) of different proportions at selected pressure-temperature combination. Results showed that the low crystalline *A. paniculata* extract powder precipitated from CO₂-Acetone system at 150 bar, 40 °C had the highest aqueous solubility of andrographolide (0.06 mg/mL), a two-fold increment than extract. However, absorption of andrographolide could still be poor for its low dissolution rate in Simulated Intestinal Fluid (SIF), pH 7.4 (0.06 mg/mL release in 90 min). Therefore, SEDS co-precipitation of *A. paniculata* with polymers were conducted using the best SEDS precipitation parameters (150 bar, 40 °C, CO₂-Acetone system). SEDS co-precipitated *A. paniculata* with Pluronic F127 or Eudragit EPO exhibited poorer andrographolide dissolution in SIF (< 0.03 mg/mL release in 90 min). SEDS co-precipitated *A. paniculata* with Eudragit L100-55 in Eudragit L100-55:*A. paniculata* mass ratio (2:25) showed improved dissolution rate of andrographolide in SIF (0.06 mg/mL release in 45 min) and mass ratio increment to 6:25 resulted in higher andrographolide release and dissolution rate (0.09 mg/mL release in 30 min). Only 20-30% of andrographolide from SEDS co-precipitates (6:25) was degraded after two-month storage whereas 30-60% of andrographolide was degraded after addition into beverages in a day. Since acetone content in SEDS co-precipitates (6:25) was < 0.1 ppm, its increased brine shrimp cytotoxicity (LC₅₀ = 46.46 µg/mL) could be attributed to higher aqueous solubility and dissolution rate of andrographolide.

Abstrak tesis yang dikemukakan kepada Senat Universiti Putra Malaysia sebagai memenuhi keperluan untuk ijazah Doktor Falsafah

SOLUTION-ENHANCED DISPERSION BY SUPERCRITICAL CARBON DIOXIDE DALAM PENGHASILAN SERBUK *Andrographis paniculata* (Burm.f.) Wall. ex Nees

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Sifat ketidaklarutan dalam air dan kadar pembubaran *andrographolide* yang lemah mengakibatkan kecenderungannya kepada kadar kemasukan yang rendah. Strategi-strategi seperti pengurangan kristalografi dan peningkatan kawasan permukaan sedia ada untuk pembubaran *andrographolide* telah dilaksanakan melalui *Solution-Enhanced Dispersion by Supercritical Carbon Dioxide (SEDS)*. Kajian dimulai dengan mengidentifikasikan trend presipitasi *SEDS* ekstrak *Andrographis paniculata* (Burm.f.) Wall. ex Nees daripada sistem karbon dioksida (CO₂)-Aseton dan sistem CO₂-Aseton:Etanol (v/v) 1:1 dalam kombinasi tekanan (100, 150 bar) dan suhu (40, 50 °C) berbeza. Perubahan kelarutan dalam air *andrographolide* selepas presipitasi turut dikaji. Modifikasi kelarutannya dalam air kemudian dilaksanakan melalui perubahan presipitasi dengan penggunaan sistem CO₂-Aseton:Etanol dalam proporsi berlainan (v/v) pada kombinasi tekanan dan suhu ditentukan. Didapati bahawa serbuk ekstrak *A. paniculata* dipresipitaskan daripada sistem CO₂-Aseton pada 150 bar, 40 °C mengalami pengkristalan rendah dan mempunyai kelarutan dalam air *andrographolide* yang tertinggi (0.06 mg/mL), peningkatan berganda berbanding ekstrak. Namun, penyerapan *andrographolide* akan disekat jika kadar pembubarannya masih rendah dalam persekitaran *Simulated Intestinal Fluid (SIF)*, pH 7.4 (0.06 mg/mL pembubaran dalam 90 min). Oleh itu, kopresipitasi *SEDS* ekstrak dengan polimer telah dilaksanakan berdasarkan parameter presipitasi *SEDS* terbaik (150 bar, 40 °C, sistem CO₂-Aseton). Kopresipitasi *SEDS* ekstrak dengan Pluronic F127 mahupun Eudragit EPO melemahkan pembubaran *andrographolide* dalam persekitaran *SIF* (< 0.03 mg/mL pembubaran dalam 90 min). Kopresipitasi *SEDS* dengan nisbah jisim Eudragit L100-55 dan ekstrak (2:25) meningkatkan kadar pembubaran *andrographolide* dalam persekitaran *SIF* (0.06 mg/mL pembubaran dalam 45 min). Peningkatan nisbah jisim hingga 6:25 turut meningkatkan pembubaran dan kadar pembubaran *andrographolide* dalam persekitaran *SIF* (0.09 mg/mL pembubaran dalam 30 min). Hanya 20-30% *andrographolide* daripada serbuk kopresipitasi *SEDS* (6:25) mengalami degradasi selepas dua bulan penyimpanan malah 30-60% *andrographolide* mengalami degradasi selepas sehari dimasukkan dalam minuman. Oleh sebab kandungan sisa aseton serbuk kopresipitasi *SEDS* yang rendah (<

0.1 ppm), peningkatan aktiviti sitotoksiknya terhadap udang air garam ($LC_{50} = 46.46 \mu\text{g/mL}$) berkemungkinan disebabkan oleh peningkatan kelarutan dan kadar pembubaran *andrographolide*.



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This thesis was submitted to the Senate of Universiti Putra Malaysia and has been accepted as fulfillment of the requirement for the degree of Doctor of Philosophy. The members of the Supervisory Committee were as follows:

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LIST OF ABBREVIATIONS

AF	Amount of Andrographolide in SEDS Feed
AG	Andrographolide
AR	Andrographolide Recovery
ARISE	Atomized Rapid Injection for Solvent Extraction
AS	Amount of Andrographolide in SEDS Precipitated <i>A. paniculata</i> powder
ASES	Aerosol Solvent Extraction System
AUC	Area under the Curve (Plasma Concentration)
BCS	Biopharmaceutics Classification System
C_{\max}	Maximum Concentration
$C_{\max, \text{pH } 1.2}$	Maximum Concentration in Simulated Gastric Fluid, pH 1.2
$C_{\max, \text{pH } 7.4}$	Maximum Concentration in Simulated Intestinal Fluid, pH 7.4
CAN-BD	Carbon Dioxide Assisted-Nebulization with a Bubble Dryer
DELOS	Depressurization of an Expanded Liquid Organic Solution
DSC	Differential Scanning Calorimetry
Ee	Encapsulation Efficiency
FT-IR	Fourier Transform Infrared
GAMA	Gas-Assisted Melting Atomization
GAS	Gas Anti-Solvent
GIT	Gastrointestinal
ICH	International Conference on Harmonization
LC ₅₀	Median Lethal Concentration
LD ₅₀	Median Lethal Dose
MCP	Mixture Critical Point

MLNE	Multilayered Nanoemulsion
MRT	Mean Residence Time
P_c	Critical Pressue
$P_{c,mix}$	Critical Pressure of Mixture
PBS	Phosphate Buffered Saline
PCA	Particles by Compressed Anti-Solvent
PDE	Permitted Daily Exposure
PGSS	Particles from Gas-Saturated Solution
PGSS-DRYING	Particles from Gas-Saturated Solutions-Drying
PID	Proportional-Integral-Derivative
PPRGEL	Pressure Reduction of Gas-Expanded Liquids
Q_{5min}	Cumulative Release (5 min)
RESOLV	Rapid Expansion of Supercritical Solution into a Liquid Solvent
RESS	Rapid Expansion of Supercritical Fluids
RESSAS	Rapid Expansion of Supercritical Solution into an Aqueous Solution
RESS-N	Rapid Expansion of Supercritical Solution with a Non-Solvent
RH	Relative Humidity
RTD	Ready-To-Drink
SAA	Supercritical Assisted Atomization
SAILA	Supercritical Assisted Injection in a Liquid Anti-Solvent
SAS	Supercritical Anti-Solvent
SASD	Supercritical CO ₂ -Assisted Spray Drying
SAS-EM	Supercritical Anti-Solvent Precipitation with Enhanced Mass Transfer

sc-CO ₂	Supercritical Carbon Dioxide
SCFs	Supercritical Fluids
SEDS	Solution-Enhanced Dispersion by Supercritical Carbon Dioxide
SEM	Scanning Electron Microscopy
SFEE	Supercritical Fluid Extraction of Emulsions
SGF	Simulated Gastric Fluid
SIF	Simulated Intestinal Fluid
SLF	Simulated Lung Fluid
SLN	Solid Lipid Nanoparticle
SMEDDS	Self-Microemulsifying Drug Delivery System
T _c	Critical Temperature
T _g	Glass Transition Temperature
T _m	Melting Point
T _{max}	Time to Reach Maximum Concentration
t _{1/2}	Half-life
t _{75%}	Time (75% Cumulative Release)
t _{90%}	Time to Decay to 90% of Its Original Concentration (Shelf Life)
TGA	Thermogravimetric
VLE	Vapor-Liquid Equilibrium
XRD	X-Ray Diffraction

CHAPTER 1

INTRODUCTION

1.1 Background

Andrographis paniculata (Burm.f.) Wall. ex Nees is a perennial herb that belongs to the family of Acanthaceae. Locally, it is named as “Hempedu Bumi”. It is also called as “King of bitters” due to its extremely characteristic bitter taste. It grows widely in tropical area of South East Asia, China, and India for traditional medicine purpose to help against fever, dysentery, diarrhea, inflammation, and sore throat (Wongkittipong et al., 2004).

According to Sareer et al. (2012), more than 80 compounds extracted and isolated from this plant had been reported. Type of compounds ranged from diterpene lactones, to a number of polyphenols such as flavonoids, xanthones, quinic acid, and noriridoids (Xu et al., 2012; Xu and Wang, 2011). The characteristic bitter taste as well as the bioactive properties of this plant are due to presence of diterpene lactones with the most abundant one – andrographolide contributing to most of the plant’s bitterness and bioactivities (Sareer et al., 2014).

Notwithstanding, andrographolide of which belongs to Class II of the Biopharmaceutics Classification System (BCS) is a poorly soluble and highly permeable drug (Zhang et al., 2016). Its high permeability to intestinal mucosa is attributed to its hydrophobicity. However, it must first dissolve in gastrointestinal fluids before it permeates intestinal membrane to reach systemic circulation (Miller et al., 2007). The rate limiting step for absorption of Class II drug such as andrographolide will be its solubility and dissolution (Savjani et al., 2012). Common strategies that are used to address low drug solubility are categorized as (a) physical modifications, (b) chemical modifications, and (c) miscellaneous as suggested (Savjani et al., 2012). While certain strategies employed during addressing low drug solubility might result in only either physical or chemical modifications of drug, strategy involves usage of supercritical fluids (SCFs) technology could be able to cater both physical and chemical modifications of poor soluble drug as Yasuji et al. (2008) had suggested that solubility improvement of poor aqueous soluble drug by SCFs was due to drug micronization (physical modification) and composite particle formation (chemical modification), all achievable in one step process.

1.2 Problem Statement

Despite of the therapeutic functions of andrographolide, it has been categorized by the BCS as a poorly soluble and highly permeable drug (Zhang et al., 2016). Solubility is an important parameter to allow the achievement of desired active concentration in systemic circulation for a targeted response, be it a pharmacological response for a drug (Savjani et al., 2012). Poorly soluble drug such as andrographolide will therefore need higher dosage or frequent dosage in order to reach the targeted concentration for a response. In

order to improve its absorption, the aqueous solubility of andrographolide needs to be improved so that it can present in the form of solution (aqueous gastrointestinal fluids) at the site of absorption to be absorbed. However, with improved aqueous solubility, the absorption of andrographolide can still be poor for its low dissolution rate. Andrographolide needs to be readily release from dosage while at the same time to be soluble in fluids in order for absorption. Therefore poor aqueous solubility and poor dissolution rate of andrographolide are both the limiting features for its low oral bioavailability that need to be solved.

1.3 Objectives

- (1) To investigate the effect of organic solvents and their mixture at different pressure-temperature combination applied during Solution-Enhanced Dispersion by Supercritical Carbon Dioxide (SEDS) precipitation of *A. paniculata* extract on aqueous solubility of andrographolide.
- (2) To investigate the effect of polymers and their core/polymer mass ratio applied in SEDS co-precipitation on dissolution of andrographolide.
- (3) To characterize the SEDS co-precipitates with the best andrographolide dissolution profile in Simulated Intestinal Fluids (SIF), pH 7.4 in terms of andrographolide content, encapsulation efficiency, morphology, crystallinity, thermal stability, core-polymer interaction.
- (4) To evaluate the storage stability profile of andrographolide in SEDS co-precipitates after being stressed under different storage temperature.
- (5) To evaluate the application stability of andrographolide in SEDS co-precipitates after being added into different Ready-To-Drink (RTD) beverages of different pH, respectively.
- (6) To identify the residual solvent content of SEDS co-precipitates and in vivo toxicity of the SEDS co-precipitates in comparison with crude *A. paniculata* extract and SEDS precipitated *A. paniculata* extract powder.

1.4 Thesis Outline

The focus of this study is on improvement of aqueous solubility and dissolution of andrographolide from *A. paniculata* extract using SEDS approach. The rationale of study is as a form of particle engineering to micronize *A. paniculata* extract as well as to co-precipitate *A. paniculata* extract with polymer in order to improve aqueous solubility and dissolution of andrographolide under the target medium. This helps to unlock the therapeutic potential of the medicinal plant - *A. paniculata* as the bioavailability of andrographolide from this plant is often restricted by its poor aqueous solubility and poor dissolution.

The first working chapter focuses on micronization (precipitation) of sticky lump of *A. paniculata* extract into *A. paniculata* powder form using SEDS approach. It involves the study of effect of organic solvents and their mixture at different pressure-temperature combination applied during SEDS precipitation on aqueous solubility of andrographolide. The precipitation of andrographolide from *A. paniculata* extract can be

modified in view of the different solubility of anti-solvent in solvent that may result in a change of precipitation as well as physical and chemical properties of powder precipitated. The objective is to improve aqueous solubility of andrographolide by increasing surface area of particle as well as by reducing crystallinity of andrographolide precipitated. The type of solvent, SEDS operating pressure and temperature that resulted in precipitation of powder with the highest aqueous solubility of andrographolide will then be applied in the second working chapter.

Co-precipitation of *A. paniculata* extract with polymer is required to unlock the dissolution improvement of andrographolide after the improvement of aqueous solubility of andrographolide through SEDS micronization as shown in the previous chapter. The second working chapter involves study of co-precipitation of *A. paniculata* extract with three different polymers (Pluronic F127, Eudragit EPO, Eudragit L100-55) with the objective to improve dissolution of andrographolide in the target medium by reducing crystallinity of andrographolide. Polymer that can target improved dissolution of andrographolide in SIF, pH 7.4 will be selected for co-precipitation with *A. paniculata* extract in different mass ratios in order to evaluate its concentration effect on andrographolide dissolution in that target environment. Co-precipitates with the highest andrographolide dissolution in SIF, pH 7.4 will be further characterized in terms of its andrographolide content, encapsulation efficiency, thermal and chemical properties.

In the third working chapter, the stability of the andrographolide from the SEDS co-precipitates will be further tested under different stressed storage conditions for two months. The stability of andrographolide from SEDS co-precipitates will also be evaluated after their addition into different pH RTD beverages for three days. This part of study represents the application stability test of SEDS co-precipitates which can provide an insight on how the degradation of andrographolide will vary with time under the presence of food matrix since its targeted application is a food applicable herbal powder (source of andrographolide) for enrichment purpose. The amount of andrographolide remained will be quantified as a way to evaluate the stability of andrographolide from SEDS co-precipitates. Toxicity of the SEDS co-precipitates will be further tested using brine shrimp as the model to identify the lethal concentration (LC_{50}) after the solvent residue test on the SEDS co-precipitates. A study overview is shown in Table 1.1.

Table 1.1: Study overview

	Stage	Rationale	Outcome
1 st	Preparation and deposition of voucher specimen (Voucher No. SK 2767/15) in herbarium	Authentication by botanist	Authenticated whole <i>A. paniculata</i> plant
2 nd	Washing, draining, cutting, drying, grinding	Pre-treatment before extraction	Dried ground whole <i>A. paniculata</i> plant powder
3 rd	Conventional solid-liquid extraction (Cold Maceration)	Extract obtained as raw material of study	Sticky crude <i>A. paniculata</i> extract
4 th	SEDS precipitation of extract	Study precipitation pattern of <i>A. paniculata</i> extract, Modification of precipitation (solvent/solvent mixture at different P-T) to improve aqueous solubility	SEDS precipitated <i>A. paniculata</i> extract powder with improved aqueous solubility of andrographolide
5 th	SEDS co-precipitation of extract with polymers	Improvement of dissolution of andrographolide with aid of polymers	SEDS co-precipitates with improved dissolution of andrographolide in target medium
6 th	Stability study of SEDS co-precipitates	Storage stability at different temperature, Application stability (after addition of powder into different RTD beverages with different pH	Stable SEDS co-precipitates
7 th	Toxicity study	Solvent residue check on SEDS co-precipitates, Brine shrimp toxicity assay for LC ₅₀ determination	SEDS co-precipitates with negligible solvent content (< 0.1 ppm) and LC ₅₀ = 46.46 µg/mL



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