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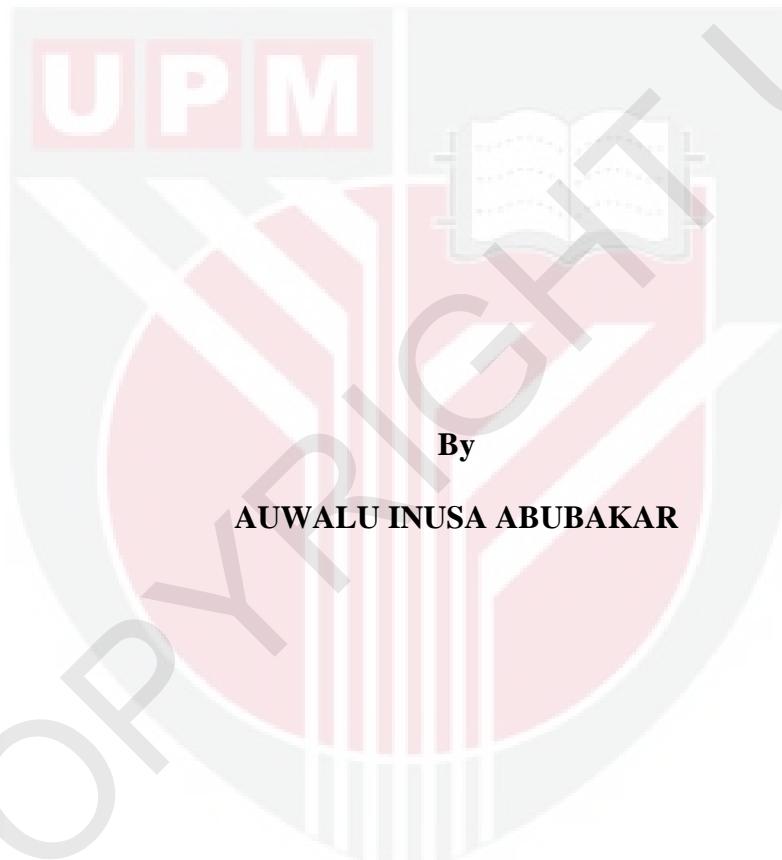
***STRUCTURAL, OPTICAL AND THERMAL PROPERTIES OF  
Sm<sub>2</sub>O<sub>3</sub>/Bi<sub>2</sub>O<sub>3</sub>-DOPED ZINC SILICATE GLASS CERAMICS FROM RICE  
HUSK ASH***

AUWALU INUSA ABUBAKAR

FS 2017 49



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HUSK ASH**



Thesis submitted to the School of Graduate Studies, Universiti Putra Malaysia,  
in Fulfillment of the Requirements for the Degree of Doctor of Philosophy

July 2017

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## **DEDICATION**

To my beloved parents Alhaji Inusa Abubakar and Malama Maryam Ahmad Gwadabe

For their absolute love and support

To my distinguish brothers, sisters and entire family

For making my life comfortable

To my dazzling wife Zainab Abdullahi Mukhtar

For her ultimate love and care

To all my cherished children

Yasin, Maryam, Almusari, Kanzillahi, Fatima and Suhaila

For their outstanding prayers

To all my humble friends

For making my life full of joy and happiness

To all my noble lecturers

For helping me in way or other throughout this remarkable journey

Thank you all

Abstract of thesis presented to the Senate of Universiti Putra Malaysia in fulfillment  
of the requirement for the degree of Doctor of Philosophy

**STRUCTURAL, OPTICAL AND THERMAL PROPERTIES OF  
 $\text{Sm}_2\text{O}_3/\text{Bi}_2\text{O}_3$ -DOPED ZINC SILICATE GLASS CERAMICS FROM RICE  
HUSK ASH**

By

**AUWALU INUSA ABUBAKAR**

**July 2017**

**Chairman : Associate Professor Halimah Mohammed Kamari, PhD**  
**Faculty : Science**

The solid waste disposal has become a persistence problem in our communities. This evolves due to our rapid increase in population that leads to the growth of our industrial and agricultural sectors. A dramatic increase of solid waste deposition is experienced. Virtually, different collection of items is disposed, which change the compositional statues of our environments. Formally, solid waste disposal was carried out by incinerators technique. However, the burning of solid waste materials pollutes the surroundings. The community screams against the hazardous air pollution from incineration of waste materials. Hence, the technique changes to landfills solid waste disposal. Besides, this process tremendous number of problems is encompassing it, which changes the scientists and researchers way of thinking to the conversion of waste materials. Therefore, a subset of agricultural waste known as waste rice husk is converted to a useful material, known as white rice husk ash (WRHA) is used in carrying out this experimental work for this thesis. Conventional melt quenching technique was used to synthesis and characterize  $\text{Sm}^{3+}/\text{Bi}^{3+}$  doped zinc silicate derived from WRHA for the structural, optical and thermal properties. Some of the major findings were given as followed. At high sintering temperature, Porosities were decreased due to the diffusion of the ions into based precursor. The poly grains turned to aggregate with one another with increasing doping concentrations. The phonon was responsible for the conveying the heat energy. An increasing heat movement was noticed with growing samarium dopant concentrations. However, a fluctuated heat movement was observed due to the oxidation and de oxidation of  $\text{Bi}^{3+}$  ions. The decrement energy band gaps of the samples were occurring because of the conversion of bridging oxygen to non-bridging oxygen. Hence, in conclusion Samarium/Bismuth oxides doped zinc silicates were successfully synthesized via melt quenching technique. Porosity decreases with increase of  $\text{Sm}^{3+}/\text{Bi}^{3+}$  ions concentrations. The poly grains increase in size with increasing of  $\text{Sm}^{3+}/\text{Bi}^{3+}$  ions concentrations. The sample agglomerated more with 5 wt% increment of  $\text{Sm}^{3+}/\text{Bi}^{3+}$  ions concentrations. Thermal diffusivity increases with increase of samarium concentrations. Thermal diffusivity fluctuates with increase of bismuth concentrations. Optical band gap of

bismuth doped zinc silicate decreases with increase of bismuth concentrations. Optical band gap of samarium doped zinc silicate increases at 3 wt% and decreases with 5 wt% samarium concentrations. Refractive index decreases with Sm 3 wt% and increases with Sm 5 wt% while molar refraction and polarizability increase with samarium concentrations up to 3 wt% and decrease to 5wt%. Refractive index, molar refraction and polarizability decrease with increasing of bismuth concentrations. Finally, the outcomes of XRD, FESEM, and FTIR transmittance spectra showed the formation of zinc silicate phase in the precursor ZnO-WRHA glass ceramics. The FESEM micrograms indicated the formation of well-defined zinc silicate phase with highly agglomerated Samarium/Bismuth oxides -doped zinc silicate glass ceramics. The crystallite size acquired from XRD pattern differs in the range of 54-59 nm for samarium oxides -doped zinc silicate and 41-94 nm for bismuth oxides -doped zinc silicate. The presence of Zn-O and Si-O-Zn vibration bands in Infrared spectra shows the features of willemite phase formation. The UV-Visible analysis displays the value of optical band gap, which reduces with growing in doping concentrations, and the sintering temperature improves the willemite crystallinity in the precursor ZnO-WRHA glass ceramics. Furthermore, the luminescence spectra of glass-ceramics establish that the  $\text{Sm}^{3+}$ /  $\text{Bi}^{3+}$  ions are diffusing into the glass-ceramic phase and through that improved the photoluminescence of the doped willemite glass-ceramics. Hence, the emission intensity of  $\text{Zn}_2\text{SiO}_4:\text{Sm}^{3+}$  phase showed red luminescence focused at 646.71 nm whereas the emission fixed at 621.57 nm developed from  $\text{Zn}_2\text{SiO}_4:\text{Bi}^{3+}$  phase matches to the  $^4\text{G}_{5/2} - ^6\text{H}_{7/2}$  transition. The thermal diffusivities values grow with dopants proportions with the least value of 0.2039 - 0.1392 mm<sup>2</sup>/s and highest value of 0.4375–0.2653 mm<sup>2</sup>/s for  $\text{Zn}_2\text{SiO}_4:\text{Sm}^{3+}$ . Again, the thermal diffusivities values fluctuate with dopant concentrations with the least value of 0.2008–0.1383 mm<sup>2</sup>/s and the highest value of 0.2329-0.1570 mm<sup>2</sup>/s for  $\text{Zn}_2\text{SiO}_4:\text{Bi}^{3+}$ . For that reason,  $\text{Zn}_2\text{SiO}_4:\text{Sm}^{3+}$  glass-ceramics seems to be favorable for use in solid-state lasers as a promising phosphor material. On the other hand,  $\text{Zn}_2\text{SiO}_4:\text{Bi}^{3+}$  glass-ceramics suggests being favorable for use on radar screen as a promising phosphor material.

Abstrak tesis yang dikemukakan kepada Senat Universiti Putra Malaysia sebagai memenuhi keperluan untuk ijazah Doktor Falsafah

**SIFAT STRUKTUR, OPTIK, DAN TERMA  $\text{Sm}_2\text{O}_3/\text{Bi}_2\text{O}_3$ -DIDOP  
SERAMIK ZINK SILIKAT DARIPADA ABU SEKAM PADI**

Oleh

**AUWALU INUSA ABUBAKAR**

**Julai 2017**

**Pengerusi : Profesor Madya Halimah Mohammed Kamari, PhD  
Fakulti : Sains**

Masalah pelupusan sisa pepejal telah menjadi satu masalah yang berterusan dalam masyarakat kita. Perkembangan hal ini disebabkan oleh peningkatan populasi yang pesat yang membawa kepada pertumbuhan sektor perindustrian dan pertanian. Pemendapan sisa pepejal telah mengalami peningkatan mendadak. Secara kasarnya, pelbagai koleksi barang yang berbeza telah dilupuskan, dimana ia mengubah komposisi pejal persekitaran kita. Secara rasmi, pelupusan sisa pepejal telah dijalankan menggunakan teknik insinerator. Walaubagaimanapun, pembakaran bahan sisa pepejal mencemarkan persekitaran. Komuniti masyarakat membangkang sekeras-kerasnya pencemaran udara berbahaya yang terhasil daripada pembakaran bahan-bahan buangan. Oleh itu, teknik ini berubah kepada tapak pelupusan sisa pepejal. Selain itu, proses ini merangkumi jumlah masalah yang besar, yang mengubah cara pemikiran para saintis dan penyelidik bagi penukaran bahan-bahan buangan. Oleh itu, satu subset sisa pertanian yang dikenali sebagai sisa sekam padi telah ditukarkan kepada bahan yang berguna, iaitu sisa abu sekam padi (WRHA) yang digunakan kerja-kerja eksperimen bagi tesis ini. Teknik pencairan pelindapkejutan secara konvensional telah digunakan bagi proses sintesis dan pencirian  $\text{Sm}^{3+}/\text{Bi}^{3+}$  yang didop zink silikat berasal daripada WRHA bagi sifat-sifat struktur, optik dan terma. Antara penemuan utama adalah seperti berikut. Pada suhu pensinteran yang tinggi, keporosan berkurang disebabkan oleh penyebaran ion ke dalam dasar pelopor. Peningkatan konsentrasi bahan dop menukar butiran poli menjadi gumpalan antara satu sama lain. Fonon bertanggungjawab untuk menyalurkan tenaga haba. Peningkatan pergerakan haba telah disedari dengan peningkatan konsentrasi bahan dop samarium. Walaubagaimanapun, pergerakan haba yang berubah-ubah telah dilihat hasil pengoksidaan dan anti pengoksidaan  $\text{Bi}^{3+}$  ion. Penyusutan tenaga jurang jalur dalam sampel telah berlaku disebabkan oleh penukaran oksigen dengan penitian kepada oksigen bukan penitian. Oleh itu, sebagai kesimpulan Samarium/ Bismut oksida yang didop zink silikat telah berjaya disintesis melalui teknik pencairan pelindapkejutan. Keporosan berkurang dengan peningkatan konsentrasi ion  $\text{Sm}^{3+}/\text{Bi}^{3+}$ . Saiz butiran poli meningkat dengan peningkatan konsentrasi ion  $\text{Sm}^{3+}/\text{Bi}^{3+}$ . Penggumpalan sampel

bertambah bagi peningkatan 5 wt% konsentrasi ion  $\text{Sm}^{3+}/\text{Bi}^{3+}$ . Kemeresapan haba meningkat dengan peningkatan konsentrasi samarium. Kemeresapan haba berubah-ubah dengan peningkatan konsentrasi bismut. Jurang jalur optik bagi bismut yang didop zink silikat berkurangan dengan peningkatan konsentrasi bismut. Jurang jalur optik bagi samarium yang didop zink silikat meningkat pada 3 wt% dan berkurang dengan 5 wt% konsentrasi samarium. Indeks biasan menurun dengan Sm 3 wt% dan bertambah dengan Sm 5 wt%, manakala molar pembiasan dan keterkutuban meningkat dengan kepekatan samarium sehingga 3 wt% dan berkurang untuk 5 wt%. Indeks biasan, molar pembiasan dan keterkutuban menurun dengan peningkatan konsentrasi bismut. Akhir sekali, hasil XRD, FESEM, dan pemindahan spektrum FTIR menunjukkan pembentukan fasa zink silikat di dalam pelopor seramik kaca  $\text{ZnO-WRHA}$ . Mikrogram FESEM menunjukkan pembentukan fasa zink silikat yang jelas dengan penggumpalan Samarium/ Bismut oksida -didop seramik kaca zink silikat yang tinggi. Saiz hablur yang diperolehi dari corak XRD berbeza dalam julat 54-59 nm bagi samarium oksida -didop zink silikat dan 41-94 nm bagi bismut oksida -didop zink silikat. Kehadiran jalur getaran  $\text{Zn-O}$  dan  $\text{Si-O-Zn}$  di dalam spektrum inframerah menunjukkan ciri-ciri pembentukan fasa wilemit. Analisis UV-Visible memaparkan nilai bagi jurang jalur optik, dimana ia berkurang dengan peningkatan konsentrasi bahan dop, dan suhu pensinteran meningkatkan penghabluran wilemit di dalam pelopor seramik kaca  $\text{ZnO-WRHA}$ . Tambahan pula, spektrum luminesens bagi seramik-kaca membuktikan bahawa  $\text{Sm}^{3+}/\text{Bi}^{3+}$  ion meresap ke dalam fasa seramik-kaca dan meningkatkan fotoluminesens pendopkan wilemit seramik-kaca. Oleh itu, intensiti pelepasan bagi fasa  $\text{Zn}_2\text{SiO}_4$ :  $\text{Sm}^{3+}$  menunjukkan luminesens merah tertumpu di 646.71 nm manakala pelepasan tetap pada 621.57 nm yang dikembangkan dari fasa  $\text{Zn}_2\text{SiO}_4$ :  $\text{Bi}^{3+}$  berpadanan dengan transisi  ${}^4\text{G}_{5/2} - {}^6\text{H}_{7/2}$ . Nilai peyerapan haba berkembang dengan perkadaran bahan dop dimana nilai paling sedikit adalah 0.2039-0.1392  $\text{mm}^2/\text{s}$  dan nilai tertinggi adalah 0.4375-0.2653  $\text{mm}^2/\text{s}$  untuk  $\text{Zn}_2\text{SiO}_4$ :  $\text{Sm}^{3+}$ . Sekali lagi, nilai peyerapan haba berubah-ubah dengan konsentrasi pendopan dengan nilai paling sedikit pada 0.2008-0.1383  $\text{mm}^2/\text{s}$  dan nilai tertinggi pada 0.2329-0.1570  $\text{mm}^2/\text{s}$  bagi  $\text{Zn}_2\text{SiO}_4$ :  $\text{Bi}^{3+}$ . Oleh kerana itu,  $\text{Zn}_2\text{SiO}_4$ :  $\text{Sm}^{3+}$  seramik-kaca sesuai untuk digunakan di dalam laser keadaan pepejal sebagai bahan fosfor yang mempunyai harapan yang cerah. Sebaliknya,  $\text{Zn}_2\text{SiO}_4$ :  $\text{Bi}^{3+}$  seramik-kaca adalah dicadangkan sesuai bagi penggunaan skrin radar sebagai bahan fosfor mempunyai harapan yang cerah.

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May Allah (S.W.A.) bless all the people that made this research work a success ‘Amin’.

I certify that a Thesis Examination Committee has met on 14 July 2017 to conduct the final examination of Auwalu Inusa Abubakar on his thesis entitled "Structural, Optical and Thermal Properties of Sm<sub>2</sub>O<sub>3</sub>/Bi<sub>2</sub>O<sub>3</sub>-Doped Zinc Silicate Glass Ceramics Derived from Rice Husk Ash" in accordance with the Universities and University Colleges Act 1971 and the Constitution of the Universiti Putra Malaysia [P.U.(A) 106] 15 March 1998. The Committee recommends that the student be awarded the Doctor of Philosophy.

Members of the Thesis Examination Committee were as follows:

**Sidek bin Hj. Ab Aziz, PhD**

Professor

Faculty of Science  
Universiti Putra Malaysia  
(Chairman)

**Abdul Halim bin Shaari, PhD**

Professor

Faculty of Science  
Universiti Putra Malaysia  
(Internal Examiner)

**Khamirul Amin bin Matori, PhD**

Associate Professor

Faculty of Science  
Universiti Putra Malaysia  
(Internal Examiner)

**Yasser Bakr Saddeek Muhammed, PhD**

Professor

Al Azhar University  
Egypt  
(External Examiner)



**NORAINI AB. SHUKOR, PhD**

Professor and Deputy Dean  
School of Graduate Studies  
Universiti Putra Malaysia

Date: 4 September 2017

This thesis was submitted to the Senate of Universiti Putra Malaysia and has been accepted as fulfillment of the requirement for the degree of Doctor of Philosophy. The members of the Supervisory Committee were as follows:

**Halimah Mohammed Kamari, PhD**

Associate Professor

Faculty of Science

Universiti Putra Malaysia

(Chairman)

**Chan Kar Tim, PhD**

Senior Lecturer

Faculty of Science

Universiti Putra Malaysia

(Member)

**Nizam Tamchek, PhD**

Senior Lecturer

Faculty of Science

Universiti Putra Malaysia

(Member)

---

**ROBIAH BINTI YUNUS, PhD**

Professor and Dean

School of Graduate Studies

Universiti Putra Malaysia

Date:

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Name of  
Chairman of  
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---

Assoc. Prof. Dr. Halimah Mohammed Kamari

---

Signature:

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Member of  
Supervisory  
Committee:



---

Dr. Chan Kar Tim

---

Signature:

Name of  
Member of  
Supervisory  
Committee:

---

Dr. Nizam Tamchek

---

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## LIST OF ABBREVIATIONS

Zn <sub>2</sub> SiO <sub>4</sub>	Willemite
Zn <sub>2</sub> SiO <sub>4</sub> :Sm <sup>3+</sup>	Samarium doped Willemite
Zn <sub>2</sub> SiO <sub>4</sub> : Bi <sup>3+</sup>	Bismuth doped Willemite
Sm	Samarium
Bi	Bismuth
WRH	Waste rice husk
WRHA	White rice husk ash
SiO <sub>2</sub>	Silicon dioxide
SiO <sub>3</sub>	Silicon trioxide
ZnO	Zinc oxide
CaO	Calcium oxide
K <sub>2</sub> O	Potassium oxide
Fe <sub>2</sub> O <sub>3</sub>	Ferric oxide
MnO	Manganese oxide
Uv	Ultraviolet
XRD	X-Ray diffraction
FTIR	Fourier transform infrared
FESEM	Field emission scanning electron microscopy
EDXRF	Energy dispersive X-ray fluorescence
TEM	Transmission electron microscopy
UV-Vis	Ultraviolet-Visible
E <sub>opt</sub>	Optical band gap
PL	Photoluminescence
PVA	Polyvinyl alcohol
R <sub>m</sub>	Molar refraction
V <sub>m</sub>	Molar volume
α <sub>m</sub>	Polarizability
ρ	Density
n <sub>o</sub>	Refractive index

# CHAPTER 1

## INTRODUCTION

### 1.1 Research background

In recent years, solid waste disposal has become a persistence problem in our communities. This evolves due to our rapid increase in population that leads to the growth of our industrial and agricultural sectors. A dramatic increase of solid waste deposition is experienced. Virtually, different collection of items is disposed, which change the compositional statues of our environments. Formally, solid waste disposal was carried out by incinerators technique. However, the burning of solid waste materials pollutes the surroundings. The community screams against the hazardous air pollution from incineration of waste materials (Sembiring and Nitivattananon 2010). Hence, the technique changes to landfills solid waste disposal. Besides, this process tremendous number of problems is encompassing it, which changes the scientists and researchers way of thinking to the conversion of waste materials.



**Figure 1.1: General solid wastes** (Source: goo.gl/paFnx6)

Today in Malaysia, we can classify solids waste in to three categories domestic, industrial and agricultural wastes as shown in Figure 1.1. By using proper technique in treating these solid waste, some of them can be converted to be useful material example waste rice husk (WRH), as shown in Figure 1.3 which is a subset of agriculture solid wastes.



**Figure 1.2: Agricultural solid wastes** (<https://goo.gl/TdRmPq>)

The Figure 1.2 represents agricultural solid wastes. Waste rice husk can be converted to a useful material, known as white rice husk ash (WRHA) which is the main material used in carrying out this experimental work. White rice husk ash (WRHA) is one form of amorphous occurred at 700 °C. WRHA simply consists of silicon dioxides ( $\text{SiO}_2$ ) as major component with other slight constituents such as  $\text{SiO}_3$ ,  $\text{CaO}$ ,  $\text{K}_2\text{O}$ ,  $\text{Fe}_2\text{O}_3$ ,  $\text{MnO}$ , and  $\text{CuO}$ .



**Figure 1.3: Waste rice husks** (Source: [goo.gl/P8zPmK](https://goo.gl/P8zPmK))

WRHA can be used to make glass ceramic which is extremely inexpensive, reasonably hard, and chemically stable. Besides, WRHA glass ceramics also have excellent insulating properties and high mechanical properties. They are recognized as the radiation-sensitive dosimeter, especially glass ceramics doped with rare earth ions or

transition metal ions. WRHA glass ceramics are attractive materials for the manufacturing of low-cost computer memory, catalytic converters, cell phones, and fluorescent (Noushad et al., 2012). In this thesis we focus on zinc oxides mixed with silica obtained from white rice husk ash to form zinc silicate willemite (glass ceramic) where later doped with samarium and bismuth oxides independently (Deepa et al., 2016). Zinc oxide is a semiconductor with an extensive energy band gap and having significant binding energy. In a glass ceramic network system, zinc oxide acts as a modifier with comprehensive applications. The purpose of this research is to explore the effect of samarium and bismuth oxides variations on the structural, optical and thermal diffusivity of zinc silicate (El Ghoul et al., 2012). An accurate replacement of zinc oxide by the samarium/bismuth oxide in the zinc silicate constituents will draw toward to form a better samarium/bismuth zinc silicate glass ceramics with a broad range of uses in optical materials for various technological and scientific applications. To the best of my knowledge,  $\text{Sm}_2\text{O}_3/\text{Bi}_2\text{O}_3$  doped zinc silicate glass ceramic using white rice husk ash has not been done so far, which was additional motivation for this research work. Silicate and lead oxides are the illustrations of conventional glass ceramic formers while bismuth oxide is not, because of the small field strength of  $\text{Bi}^{3+}$  ions and high polarizability. Nevertheless, the bismuth oxide shows the most complicated function in glass ceramic formation. Bismuth oxide has an important quality of broad transmission range, nonpoisonous and high refractive index in the glass composition. The amorphous forms of silicate have extensive industrial applications. The bismuth-silicate glass ceramics have distinctive uses as optical amplifiers, third-order nonlinear susceptibility, low loss optical fibers, oscillators and infrared-transmitting materials. These glass ceramics also have high refractive index, high dielectric constant, high density, transmission in central of infrared region. The bismuth-silicate glass ceramics have vital benefits especially in sealing glass ceramics for metals, plasma panel displays, reflecting windows, thick film conductors, in the field of glass ceramics and electronic devices respectively. A considerable number of the studies were conducted on the structural and luminescence properties of doped zinc silicate such as europium doped zinc silicate manganese doped zinc orthosilicate and cerium doped zinc silicate glass ceramic. It is found that doped zinc silicate are excellent candidates for numerous industrial applications.

## 1.2 Problems statement

Within the series of conventional glass ceramics, WRHA glass ceramics have outstanding promising application due to their excellent glass forming character correlating to the discrete conventional glass ceramics (Hashmi et al., 2013). Besides, WRHA glass ceramic is accepted due to its fine mechanical and optical properties, such as high UV transparency, good chemical stability, low nonlinear refractive index and low thermal expansion coefficient leading to the strong thermal resistance (Sahu et al., 2015). WRHA glass ceramic also has large tensile fracture strength and good durability. The feasibility to regulate the physical properties such as density and refractive index of the glass ceramic is by the alteration of the glass ceramics composition (Eevon et al., 2016). A vigorous deliberation happened in the literature regarding the effect of doping of samarium/bismuth oxides into borate and telluride glass ceramic; however, the doping of samarium/bismuth oxides into WRHA glass ceramic is very few with no authoritative answers to the essential subject matter

(Matori et al., 2009). Despite, the diverse properties of WRHA glass ceramic such as good mechanical and high insulating properties have drawn attention of many researchers' due to the variety of technological and industrial applications (Patra et al., 2004). Lack of work for this glass ceramic had been investigated especially samarium and bismuth oxides doped zinc silicate (willemite) glass ceramic from white rice husk ash respectively. Therefore, an itemized study of the structural, optical and thermal properties of samarium/bismuth doped with WRHA glass ceramics are carried out and the results of this research are offered in this thesis.

### **1.3 Research objective**

The research objectives are:

1. To synthesize of samarium and bismuth oxides, doped zinc silicate (willemite) derived from white rice husk ash.
2. To determine the optical absorption, optical band gap, refractive index, molar refraction, polarizability, excitation and emission spectra of samarium/bismuth oxides, doped zinc silicate ceramics derived from white rice husk ash.
3. To study the thermal properties of samarium/bismuth oxides zinc silicate ceramics derived from white rice husk ash.

### **1.4 Hypothesis**

Zinc silicate (willemite) is a porous compound. The doping of samarium oxide and bismuth oxide will increase the crystal size structure, by the diffusion of electrons into the host matrix, which might reduce the porosity of the willemite. Further addition of the dopant may cause the structure to agglomerate. Concerning the optical properties samarium doped zinc silicate the optical band gap may increase up to a certain level then decrease to a certain level after the dopant addition. However, bismuth oxide doping zinc silicate may reveal the decrease in the optical band gap as the dopant concentration is increasing. The thermal properties may show the increase in thermal diffusivity with the increase in dopant concentration for samarium oxide doped zinc silicate. At the same time increasing in dopant concentration of bismuth oxide doped zinc silicate may fluctuates thermal diffusivities values.

## **1.5 Scope of the study**

To achieve the objectives of the study, the scopes of the research are as follow:

1. The glass ceramic samples uses weight percentage pattern (70-X)  $Zn_230SiO_4 XSm_2O_3$  with  $X = 0, 1, 3, 5$  wt% and (60-X)  $Zn_240SiO_4 XBi_2O_3$  with  $X = 0, 1, 3$  and 5 wt% to prepare WRHA glass ceramics. By using conventional melt and quenching technique.
2. The structure of (70-X)  $Zn_230SiO_4 XSm_2O_3$  with  $X = 0, 1, 3$  and 5 wt% and (60-X)  $Zn_240SiO_4 XBi_2O_3$  with  $X = 0, 1, 3$  and 5 wt% glass ceramic measures by X-ray diffraction technique to confirm the crystalline nature of the glass ceramic samples.
3. The energy dispersive X-ray fluorescence (EDXRF) measures the composition of the glass ceramic samples.
4. The Ultra-Violet Visible (UV-Vis) helps to measure the optical band gap. The Perkin Elmer LS 55 Fluorescence spectrometer is an instrument that allows measurement of photoluminescence (PL) means excitation and emission spectra of a sample with its electronic transitions.
5. The Archimedes' technique employed to measure the densities and combined with the obtained energy band gap to calculate the refractive index, molar refraction and polarizability of glass ceramic samples.
6. The Perkin Elmer TGA7 accomplishes TGA and DSC test under nitrogen gas at 1 atm. TGA test evaluate the mass changes of a sample with temperature. Moreover, NETZSCH-LFA 457 micro-flash device with an InSb sensor performed the thermal diffusivity test.

## **1.6 Importance of the study**

Preparing of zinc willemite glass ceramic is quite expensive, and the most important component of willemite glass ceramic is silica. About 98.2% purity of the burnt white rice husk ash are silica as revealed by XRF results. Currently, Malaysia is positioned as twenty-five in the world rice producing countries (US dept. of agriculture, 2016). Most of the landfills in Malaysia are filled up with industrial and agricultural waste. In few years to come with rapid development in both agricultural and industrial sectors in Malaysia, this will increase the percentage of the waste production. To reduce landfill disposal of solid waste, conversion of waste rice husk into a useful material known as silica is a decent idea. Silica combined with the zinc oxide can form, zinc silicate (willemite) which have extensive industrial application (Sulowska et al., 2016). The research on the structural, optical and thermal properties of zinc silicate doped with rare earth and post-transitional elements are not well investigated. Therefore, additional research is significantly required to be performed in order to determine the important properties of zinc silicate doped with samarium/bismuth oxides.

## 1.7 Outline of thesis

The thesis structural arrangement is designed as follows, Chapter 1 narrates of zinc oxide and waste rice husk ash, zinc silicate (willemite) glass ceramics and Sm/Bi oxides doped zinc silicate, the objectives, the problem statements, the importance of the study and the scope of the research. The Chapter 2 corresponds with the recent and preceding works executed by other scholars of glass and glass-ceramic respectively. Chapter 3 explicitly deals with the theory of the electronics transition and the thermal movement through the entire samples. The methodological outline of study clearly states the characterizations of willemite glass-ceramic and willemite doped with samarium and bismuth oxides dopants all are in Chapter 4. Chapter 5 analyzes all the findings made in the research work such as XRD, FESEM, FTIR, EDX, TEM, UV-Vis, PL,  $T_g$ , DSC and thermal diffusivity for samarium and bismuth, oxides doped zinc silicate with different weight percentages sintered at 1000 °C. Last but not the least Chapter 6 deals with the conclusion of the entire work and suggestions on how to carryout future research works.

## REFERENCES

- Abo-Mosallam, H. A., Darwish, H., and Salman, S. M. (2010). Crystallization characteristic and properties of some zinc containing soda lime silicate glasses. *Journal of Materials Science: Materials in Electronics*, 21(9), 889-896.
- Abu-Rukah, Y., and Al-Kofahi, O. (2001). The assessment of the effect of landfill leachate on ground-water quality—a case study. El-Akader landfill site—north Jordan. *Journal of Arid Environments*, 49(3), 615-630.
- Ahmed, A., Ibrahim, A. A., and El-Shorifi, F. T. (2014). Removal of Methylene Blue From Aqueous Solutionbybivo4/Mcm-41 Nanopartical. *International Journal of Chemical and Petrochemical Technology (IJCPT)*, 1(4), 1-12.
- Akihama, Y., and Hane, K. (2012). Single and multiple optical switches that use freestanding silicon nanowire waveguide couplers. *Light: Science and Applications*, 56(4), 215-221.
- Akshatha, W. A. G. H., Raviprakash, Y., Ajithkumar, M. P., Upadhyaya, V., and Kamath, S. D. (2015). Effect of  $\text{Sm}_2\text{O}_3$  on structural and thermal properties of zinc fluoroborate glasses. *Transactions of Nonferrous Metals Society of China*, 25(4), 1185-1193.
- Al-Nidawi, A. J. A., Matori, K. A., Zakaria, A., and Zaid, M. H. M. (2017). Effect of  $\text{MnO}_2$  doped on physical, structure and optical properties of zinc silicate glasses from waste rice husk ash. *Results in Physics*, 7, 955-961.
- Amin Matori, K., MohdZaid, M. H., Quah, H. J., Abdul Aziz, S. H., Abdul Wahab, Z., and MohdGhazali, M. S. (2015). Studying the effect of zno on physical and elastic properties of ( $\text{ZnO}$ ) glasses using nondestructive ultrasonic method. *Advances in Materials Science and Engineering*, 5(13), 217-223.
- Amjad, R. J., Sahar, M. R., Ghoshal, S. K., Dousti, M. R., Riaz, S., and Tahir, B. A. (2012).Optical Investigation of  $\text{Sm}^{3+}$  Doped Zinc-Lead-Phosphate Glass. *Chinese Physics Letters*, 29(8), 087304.
- Arena, U., Mastellone, M. L., and Perugini, F. (2003). The environmental performance of alternative solid waste management options: a life cycle assessment study. *Chemical Engineering Journal*, 96(1), 207-222.
- Auwalu, I. A., Halimah, M. K., Zaidan, A. W., Chan, K. T., and Abdullahi, U. (2017). Effect of bismuth oxide on structural and thermal diffusivity of waste rice ash zinc silicate (willemite) glass ceramics. *Journal of the Australian Ceramic Society*, 1-7.
- Ayuni, J. N., Halimah, M. K., Talib, Z. A., Sidek, H. A. A., Daud, W. M., Zaidan, A. W., and Khamirul, A. M. (2011). Optical Properties of Ternary  $\text{TeO}_2\text{-B}_2\text{O}_3\text{-ZnO}$  Glass System.IOP Conference Series: Materials Science and Engineering, 17(9), 879-886.
- Azlan, M.N, M.K. Halimah, S.Z. Shafinas, W.M. Daud, H. A. A. S. (2014).influence of erbium concentration on spectroscopic.Journal.masshp.net, 22(1), 148–156.

- Babu B. C. and Buddhudu S. (2014). Emission spectra of  $Tb^{3+}$ :  $Zn_2SiO_4$  and  $Eu^{3+}$ :  $Zn_2SiO_4$  sol-gel powder phosphors. *Journal of Spectroscopy and Dynamics*, 4(5), 9.
- Babu, B. C., Rudhramadevi, B. H., and Buddhudu, S. (2014). Synthesis and optical analysis of  $Cr^{3+}$ :  $Zn_2SiO_4$  nanocomposites by sol-gel method, 13(6), 1726–1728.
- Bahari, H., Abd Aziz, S. H., Kamari, H. M., Yunus, W. M. M., and Adikan, F. R. M. (2012). The effect of bismuth on the structure and mechanical properties of  $GeO_2$ - $PbO$ - $Bi_2O_3$  ternary bulk glass system. *Journal of the Ceramic Society of Japan*, 120(1403), 280–285.
- Bakier, E. (2005). Factors affecting light energy transfer in some samarium complexes. *International Journal of Photoenergy*, 07(1), 1–8.
- Bala, R., Agarwal, A., Sanghi, S., and Khassa, S. (2014). Influence of  $SiO_2$  on the structural and dielectric properties of  $ZnO$ ·  $Bi_2O_3$ ·  $SiO_2$  glasses. *Journal of Integrated Science and Technology*, 3(1), 6-13.
- Bale, S., Rahman, S., Awasthi, A. M., and Sathe, V. (2008). Role of  $Bi_2O_3$  content on physical, optical and vibrational studies in  $Bi_2O_3$ - $ZnO$ - $Bi_2O_3$  glasses. *Journal of Alloys and Compounds*, 460(1), 699-703.
- Barsoum, M. W., El-Raghy, T., Rawn, C. J., Porter, W. D., Wang, H., Payzant, E. A. and Hubbard, C. R. 1999. Thermal Properties of  $Ti_3SiC_2$ . *Journal of Physics And Chemistry of Solids* 60:429-439.
- Basavaraj, R. B., Nagabhushana, H., Prasad, B. D., Sharma, S. C., Prashantha, S. C., and Nagabhushana, B. M. (2015). A single host white light emitting  $Zn_2SiO_4:Re^{3+}$  ( $Eu$ ,  $Dy$ ,  $Sm$ ) phosphor for LED applications. *Optik-International Journal for Light and Electron Optics*, 126(19), 1745-1756.
- Batayneh, M., Marie, I., and Asi, I. (2007). Use of selected waste materials in concrete mixes. *Waste management*, 27(12), 1870-1876.
- Bateni, N. H., Hamidon, M. N., and Matori, K. A. (2014). Effect of soda-lime-silica glass addition on the physical properties of ceramic obtained from white rice husk ash. *Journal of the ceramic society of Japan*, 122(1422), 161-165.
- Bazant, Z. P. (1998). Size effect in tensile and compression fracture of concrete structures: computational modeling and design. *Aedification Publishers, Fracture Mechanics of Concrete Structures*, 3, 1905-1922.
- Bernstein, J. A., Khodursky, A. B., Lin, P. H., Lin-Chao, S., and Cohen, S. N. (2002). Global analysis of mRNA decay and abundance in *Escherichia coli* at single-gene resolution using two-color fluorescent DNA microarrays. *Proceedings of the National Academy of Sciences*, 99(15), 9697-9702.
- Besmann, T. M., and Spear, K. E. (2002). Thermochemical modeling of oxide glasses. *Journal of the American Ceramic Society*, 85(12), 2887-2894.
- Bian, S. W., Ma, Z., Cui, Z. M., Zhang, L. S., Niu, F., and Song, W. G. (2008). Synthesis of Micrometer-Sized Nanostructured Magnesium Oxide and Its High Catalytic Activity in the Claisen– Schmidt Condensation Reaction. *The Journal of Physical Chemistry C*, 112(39), 15602-15602.

- Boccaccini, A. R., Stumpfe, W., Taplin, D. M. R., and Ponton, C. B. (1996). Densification and crystallization of glass powder compacts during constant heating rate sintering. *Materials Science and Engineering: A*, 219(1-2), 26-31.
- Bondioli, F., Barbieri, L., Ferrari, A. M., and Manfredini, T. (2010). Characterization of rice husk ash and its recycling as quartz substitute for the production of ceramic glazes. *Journal of the American Ceramic Society*, 93(1), 121-126.
- Boni, M., and Large, D. (2003). Nonsulfide zinc mineralization in Europe: An overview. *Economic Geology*, 98(4), 715-729.
- Bunting, E. N. (1930). Phase equilibria in the system SiO<sub>2</sub>-ZnO. *Journal of the American Ceramic Society*, 13(1), 5-10.
- Carpenter, A., and Cramer, S. (1999). Mitigation of Alkali-Silica Reaction in pavement patch concrete that incorporates highly reactive fine aggregate. *Transportation Research Record: Journal of the Transportation Research Board*, (1668), 60-67.
- Casasola, R., Rincón, J. M., and Romero, M. (2012). Glass-ceramic glazes for ceramic tiles: a review. *Journal of Materials Science*, 47(2), 553-582.
- Chandrappa, G. T., Ghosh, S., and Patil, K. C. (1999). Synthesis and properties of willemite, Zn<sub>2</sub>SiO<sub>4</sub>, and M<sup>2+</sup>: Zn<sub>2</sub>SiO<sub>4</sub> (M = Co and Ni). *Journal of Materials Synthesis and Processing*, 7(5), 273-279.
- Chelouche, A., Djouadi, D., and Aksas, A. (2014). Structural characterization of SiO<sub>2</sub>/Zn<sub>2</sub>SiO<sub>4</sub>:Ce nanocomposite obtained by sol gel method. *Journal of the Association of Arab Universities for Basic and Applied Sciences*, 15, 10–14.
- Christensen, K. K., and Sand-Jensen, K. (1998). Precipitated iron and manganese plaques restrict root uptake of phosphorus in Lobelia dortmanna. *Canadian Journal of Botany*, 76(12), 2158-2163.
- Cimpoeşu, I., Stanciu, S., Enache, A., Rădoi, M., and Cimpoeşu, R. (1962). Differential scanning calorimetry of different metallic and non-metallic materials experimental results. *Ştiinţa şi Ingineria Materialelor*, 37.
- Colombelli, R., Srinivasan, K., Troccoli, M., Painter, O., Gmachl, C. F., Tennant, D. M., and Capasso, F. (2003). Quantum cascade surface-emitting photonic crystal laser. *Science*, 302(5649), 1374-1377.
- Corinaldesi, V., Gnappi, G., Moriconi, G., and Montenero, A. (2005). Reuse of ground waste glass as aggregate for mortars. *Waste Management*, 25(2), 197-201.
- Cui, J., Yang, Y., Hu, Y., and Li, F. (2015). Rice husk based porous carbon loaded with silver nanoparticles by a simple and cost-effective approach and their antibacterial activity. *Journal of Colloid and Interface Science*, 455, 117–124.
- Davis, J.W. and Haasz, A. A. 1998. Thermal Transport in CKC Tib<sub>2</sub>-Doped Graphite. *Journal of Nuclear Materials* 252:150-155.
- Davis, E. A. and Mott, N. F., (2012). Electronic processes in non-crystalline materials. OUP Oxford.
- Decker, R., Wang, Y., Brar, V. W., Regan, W., Tsai, H. Z., Wu, Q., and Crommie, M. F. (2011). Local electronic properties of graphene on a BN substrate via scanning tunneling microscopy. *Nano letters*, 11(6), 2291-2295.

- Deepa, A. V., Priya, M., and Suresh, S. (2016). Influence of Samarium Oxide ions on structural and optical properties of borate glasses, 11(5), 57–63.
- Ducman, V., Mladenovič, A., and Šuput, J. S. (2002). Lightweight aggregate based on waste glass and its alkali–silica reactivity. Cement and Concrete Research, 32(2), 223-226.
- Eevon, C., Halimah, M. K., Zakaria, A., Azurahanim, C. A. C., Azlan, M. N., and Faznny, M. F. (2016). Results in Physics Linear and nonlinear optical properties of  $Gd^{3+}$  doped zinc borotellurite glasses for all-optical switching applications. Results in Physics, 6, 761–766.
- Effendy, N., Wahab, Z. A., Abdul Aziz, S. H., Matori, K. A., Zaid, M. H. M., and Rashid, S. S. A. (2017). Characterization and optical properties of erbium oxide doped  $ZnO$ –SLS glass for potential optical and optoelectronic materials. Materials Express, 7(1), 59-65.
- Erhart, D., and Flügel, S. (2011). Properties of Zinc Silicate Glasses and Melts. Materials Science and Engineering, 1, 312–320.
- El Ghoul, J., and El Mir, L. (2015). Photoconversion from yellow-to-green in vanadium doped zinc silicate nanophosphor material. Superlattices and Microstructures, 82, 551–558.
- El Ghoul, J., Omri, K., Alyamani, A., Barthou, C., and El Mir, L. (2013). Synthesis and luminescence of  $SiO_2/Zn_2SiO_4$  and  $SiO_2/Zn_2SiO_4:Mn$  composite with sol-gel methods. Journal of Luminescence, 138, 218-222.
- El Ghoul, J., Omri, K., El Mir, L., Barthou, C., and Alaya, S. (2012). Sol–gel synthesis and luminescent properties of  $SiO_2/Zn_2SiO_4$  and  $SiO_2/Zn_2SiO_4:V$  composite materials. Journal of Luminescence, 132(9), 2288–2292.
- El Mir, L., and Omri, K. (2014). Photoconversion from UV-to-yellow in Mn doped zinc silicate nanophosphor material. Superlattices and Microstructures, 75, 89–98.
- El Ghoul, J., and El Mir, L. (2014). Sol–gel synthesis and luminescence of undoped and Mn-doped zinc orthosilicate phosphor nanocomposites. Journal of Luminescence, 148, 82-88.
- Eraiah, B. (2006). Optical properties of samarium doped zinc-tellurite glasses. Bulletin of Materials Science, 29(4), 375-378.
- Eraiah, B., and Bhat, S. G. (2007). Optical properties of samarium doped zinc-phosphate glasses. Journal of Physics and Chemistry of Solids, 68(4), 581–585.
- Eraiah, B. (2010). Optical properties of lead-tellurite glasses doped with samarium trioxide. Bulletin of Materials Science, 33(4), 391-394.
- Eriksson, G., and Pelton, A. D. (1993). Critical evaluation and optimization of the thermodynamic properties and phase diagrams of the  $CaO-Al_2O_3$ ,  $Al_2O_3-SiO_2$ , and  $CaO-Al_2O_3-SiO_2$  systems. Metallurgical Transactions B, 24(5), 807-816.
- Fatta, D., Papadopoulos, A., and Loizidou, M. (1999). A study on the landfill leachate and its impact on the groundwater quality of the greater area. Environmental Geochemistry and Health, 21(2), 175-190.

- Fayette, S., Smith, D. S., Agnes, S. and Christian, M. 2000. Influence Of Grain Size On The Thermal Conductivity Of Tin Oxide Ceramics. *Journal of the European Ceramic Society* 20:297-302.
- Faznny, M. F., Halimah, M. K., and Azlan, M. N. (2016). Effect of lanthanum oxide on optical properties of zinc borotellurite glass system. *Journal of optoelectronics and biomedical materials*, 8(2), 49–59.
- Feldmann, C., Jüstel, T., Ronda, C. R., and Schmidt, P. J. (2003). Inorganic luminescent materials: 100 Years of research and application. *Advanced Functional Materials*, 13(7), 511–516.
- Flyhammar, P. (1995). Analysis of the cadmium flux in Sweden with special emphasis on landfill leachate. *Journal of Environmental Quality*, 24(4), 612-621.
- Fonda, G. R. (1940). The yellow and red zinc silicate phosphors. *The Journal of Physical Chemistry*, 44(7), 851-861.
- Forés, A., Llusrà, M., Badenes, J. A., Calbo, J., Tena, M. a., and Monrós, G. (2000). Cobalt minimisation in willemite ( $\text{Co}_x\text{Zn}_{2-x}\text{SiO}_4$ ) ceramic pigments. *Green Chemistry*, 2(3), 93–100.
- Forrest, J. A., Dalnoki-veress, K., Stevens, J. R., and Dutcher, J. R. (1996). Effect of Free Surfaces on the Glass Transition Temperature of Thin Polymer Films Phys. Rev .Physical Review Letters, 77(10), 2002 – 2006.
- Fukumi, K., Chayahara, A., Kadono, K., Sakaguchi, T., Horino, Y., Miya, M., and Satou, M. (1991). Au<sup>+</sup>-ion-implanted silica glass with non-linear optical property. *Japanese journal of applied physics*, 30(4B), L742.
- Galkin, V., Kuznetsov, G., and Turkin, A. (2007). Thermal expansion of ZnSiO<sub>3</sub> high-pressure phases. *Physics and Chemistry of Minerals*, 34(6), 377-381.
- García, E., de Pablos, A., Bengoechea, M. A., Guaita, L., Osendi, M. I., and Miranzo, P. (2011). Thermal conductivity studies on ceramic floor tiles. *Ceramics International*, 37(1), 369–375.
- Gorlich, E. (1997). The effective nuclear charges and the electronegativity. Krakow Polish Academy of Art and Science, 2(1), 49–57.
- González-Borrero, P. P., Sato, F., Medina, A. N., Baesso, M. L., Bento, A. C., Baldissera, G., Ferreira Da Silva, A. (2010). Optical band-gap determination of nanostructured WO<sub>3</sub> film. *Applied Physics Letters*, 96(6), 4–6.
- Götz, J., and Masson, C. R. (1970). Trimethylsilyl derivatives for the study of silicate structures. Part I. A direct method of trimethylsilylation. *Journal of the Chemical Society A: Inorganic, Physical, Theoretical*, 2683-2686.
- Gotz, J., and Masson, C.R., (1978). *Dalton Trans. J. Chem. Soc.*, 1134-1143.
- Graça, M. P. F., da Silva, M. F., Sombra, A. S. B., and Valente, M. A. (2007). Electric and dielectric properties of a SiO<sub>2</sub>-Na<sub>2</sub>O-Nb<sub>2</sub>O<sub>5</sub> glass subject to a controlled heat-treatment process. *Physica B: Condensed Matter*, 396(1), 62-69.
- Hajer, S. S., Halimah, M. K., Azmi, Z., and Azlan, M. N. (2014). Optical properties of zinc-borotellurite doped samarium. *Chalcogenide Lett*, 11, 553-566.
- Halimah, M., and Daud, W. (2010). Optical properties of ternary tellurite glasses. *Mater. Sci. Pol.*, 28(1), 173-180.

- Hashmi, M. U., Shah, S. A., Umer, F., and Alkedy, A. S. (2013). Effect of sintering temperature on microstructure and in-vitro behavior of bioactive glass-ceramics. *Ceramics - Silikaty*, 57(4), 313–318.
- He, H., Zhang, Y., Pan, Q., Wu, G., Dong, G., and Qiu, J. (2015). Controllable synthesis of  $Zn_2GeO_4$ :Eu nanocrystals with multi-color emission for white light-emitting diodes. *J. Mater. Chem. C*, 3(21), 5419–5429.
- Hegazy, A. A., Zaher, M. M., Abd el-hafez, M. A., Morsy, A. A., and Saleh, R. A. (2010). Relation between anemia and blood levels of lead, copper, zinc and iron among children. *BMC research notes*, 3(1), 133.
- Hench, L. L. (1977). Physical chemistry of glass surfaces. *Journal of Non-Crystalline Solids*, 25(1), 343-369.
- Hitzman, M. W., Reynolds, N. A., Sangster, D. F., Allen, C. R., and Carman, C. E. (2003). Classification, genesis, and exploration guides for nonsulfide zinc deposits. *Economic Geology*, 98(4), 685-714.
- Hojamberdiev, M., Muhammedbaeva, Z., and Madhusoodana, C. D. (2009). Use of natural and thermally activated porphyrite in cement production. *Construction and Building Materials*, 23(8), 2757–2762.
- Hong, J. Y., Suh, E. H., and Kim, S. J. (2009). Context-aware systems: A literature review and classification. *Expert Systems with Applications*, 36(4), 8509-8522.
- Hooda, J., Punia, R., Kundu, R. S., Dhankhar, S., and Kishore, N. (2012). Structural and Physical Properties of  $ZnO$  Modified Bismuth Silicate Glasses. *International Scholarly Research Network Spectroscopy*, 3(12), 514-519
- Hu, A. M., Li, M., Dali, D. M., and Liang, K. M. (2005). Crystallization and properties of a spodumene-willemite glass ceramic. *Thermochimica Acta*, 437(1), 110-113.
- Hübner, P., Studer, E., Häfliger, D., Stadler, M., Wolf, C., and Looser, M. (1999). Detection of genetically modified organisms in food: critical points for quality assurance. *Accreditation and Quality Assurance: Journal for Quality, Comparability and Reliability in Chemical Measurement*, 4(7), 292-298.
- Husung, R. D., and Doremus, R. H. (1990). The infrared transmission spectra of four silicate glasses before and after exposure to water. *Journal of Materials Research*, 5(10), 2209-2217.
- Innocenzi, P., Falcaro, P., Grosso, D., and Babonneau, F. (2003). Order-disorder transitions and evolution of silica structure in self-assembled mesostructured silica films studied through FTIR spectroscopy. *The Journal of Physical Chemistry B*, 107(20), 4711-4717.
- Josephine, L.Y.C., 2003. Thermal diffusivity and electrical conductivity studies of polyaniline, polyaniline blends, polyaniline composite and ceramic samples, MsC Thesis, Universiti Putra Malaysia.
- Kang, S. J. L. (2004). Sintering: densification, grain growth and microstructure. Butterworth-Heinemann.

- Karar, A. A. (2013). Surface plasmons for enhanced metal-semiconductor-metal photodetectors. *Journal of Lightwave Technology*, 31(7), 1088-1092.
- Kavlock, R. J., Daston, G. P., DeRosa, C., Fenner-Crisp, P., Gray, L. E., Kaattari, S., and Miller, R. (1996). Research needs for the risk assessment of health and environmental effects of endocrine disruptors: a report of the US EPA-sponsored workshop. *Environmental health perspectives*, 104(Suppl 4), 715.
- Khati, V. A., Singru, R. N., and Duryodhan, D. D. (2015, January). Wealth from Waste Rice Husk. In *International Symposium on Ultrasonics* (Vol. 22, No. 24).
- Kim, E., Jiang, Z.-T., and No, K. (2000). Measurement and Calculation of Optical Band Gap of Chromium Aluminum Oxide Films. *Japanese Journal of Applied Physics*, 39(Part 1, No. 8), 4820–4825.
- Kindrat, I. I., Padlyak, B. V., and Drzewiecki, A. (2015). Luminescence properties of the Sm-doped borate glasses. *Journal of Luminescence*, 166, 264–275.
- Krsmanović, R., Antić, Ž. Marinović-Cincović, M., and Dramićanin, M. Samarium and terbium doped  $Zn_2SiO_4$  powders obtained by polymer induced sol-gel synthesis 5(1), 57-63.
- Łaczka, M., Stoch, L., and Górecki, J. (1992). Bismuth-containing glasses as materials for optoelectronics. *Journal of alloys and compounds*, 186(2), 279-291.
- Lange, F. F. (1989). Thermodynamics of densification: II, grain growth in porous compacts and relation to densification. *Journal of the American Ceramic Society*, 72(5), 735-741.
- Lenggoro, I. W., Iskandar, F., Mizushima, H., Xia, B., Okuyama, K., and Kijima, N. (2000). One-step synthesis for  $Zn_2SiO_4$ :Mn particles 0.3–1.3  $\mu m$  in size with spherical morphology and non-aggregation. *Japanese Journal of Applied Physics*, 39(10B), L1051.
- Leverenz, H. W., and Seitz, F. (1939). Luminescent materials. *Journal of Applied Physics*, 10(7), 479-493.
- Leverenz, H. W. (1949). Luminescent solids (phosphors). *Science*, 109(2826), 183-195.
- Li, Q., Creaser, D., and Sterte, J. (1999). The nucleation period for TPA-silicalite-1 crystallization determined by a two-stage varying-temperature synthesis. *Microporous and Mesoporous Materials*, 31(1), 141-150.
- Li, X., and Chen, F. (2013). Structure and luminescence properties of  $Zn_2SiO_4$ :Mn phosphor prepared with MCM-48. *Materials Research Bulletin*, 48(6), 2304-2307.
- Lin, T., Aitola, K., Novikov, S., Räsänen, M., Velagapudi, R., Sainio, J., and Khriachtchev, L. (2004). Optical and structural properties of silicon-rich silicon oxide films: Comparison of ion implantation and molecular beam deposition methods. *Physica status solidi (a)*, 208(9), 2176-2181.
- Lukić, S. R., Petrović, D. M., Dramićanin, M. D., Mitrić, M., and Đačanin, L. (2008). Optical and structural properties of  $Zn_2SiO_4$ : $Mn^{2+}$  green phosphor nanoparticles obtained by a polymer-assisted sol-gel method. *Scripta Materialia*, 58(8), 655–658.

- Lun, C., Duh, J. G., and Chiou, B. S. (2003). Low temperature sintering and crystallisation behaviour of low loss anorthite-based glass-ceramics. *Journal of Materials Science*, 38(4), 693–698.
- Machida, H., Takesue, M., and Smith, R. L. (2011). Green chemical processes with supercritical fluids: properties, materials, separations and energy. *The Journal of Supercritical Fluids*, 60, 2-15.
- Mai, M., and Feldmann, C., (2009). *Solid State Sci.* 11, 528-532.
- Makinistian, L., and Albanesi, E. A. (2007). Ab initio calculations of the electronic and optical properties of germanium selenide. *Journal of Physics. Condensed Matter*, 19, 186–211.
- Mansal, N.P. and Zhu, D.M. 2007. Effects of doping on thermal conductivity of pyrochlore oxides for advanced thermal barrier coatings. *Mater.Sci. Eng. A* 459:192-195.
- Marumo, F., and Syono, Y. A. S. U. H. I. K. O. (1971). The crystal structure of  $Zn_2SiO_4$ -II, a high-pressure phase of willemite. *Acta Crystallographica Section B: Structural Crystallography and Crystal Chemistry*, 27(10), 1868-1870.
- Marzouk, M. A., Hamdy, Y. M., El Batal, H. A., and ElDin, F. M. (2015). Photoluminescence and spectroscopic dependence of fluorophosphate glasses on samarium ions concentration and the induced defects by gamma irradiation. *Journal of Luminescence*, 166, 295–303.
- Matori, K. A., Haslinawati, M. M., Wahab, Z. A., and Ban, T. K. (2009). Producing Amorphous White Silica from Rice Husk. *Masaum Journal of Basic and Applied Sciences*, 1(3), 512–515.
- Matori, K. A., Zaid, M. H. M., Sidek, H. A. A., Halimah, M. K., Wahab, Z. A., and Sabri, M. G. M. (2010). Influence of  $ZnO$  on the ultrasonic velocity and elastic moduli of soda lime silicate glasses. *International Journal of Physical Sciences*, 5(14), 2212-2216
- McDougall, F. R., White, P. R., Franke, M., and Hindle, P. (2008). Integrated solid waste management: a life cycle inventory. John Wiley and Sons.
- Meng, X. G., Qiu, J. R., Peng, M. Y., Chen, D. P., Zhao, Q. Z., Jiang, X. W., and Zhu, C. S. (2005). Near infrared broadband emission of bismuth-doped aluminophosphate glass. *Optics Express*, 13(5), 1628-1634.
- Merkulov, A.G., and Khristoforov, B.S. (1969). *Izv. Sib. Otd. Akad. Nauk SSSR*, 153-155.
- Milde, M., Dembski, S., Osset, A., Batentschuk, M., Winnacker, A., and Sextl, G. (2014). Polymer-assisted sol-gel process for the preparation of photostimulable core/shell structured  $SiO_2/Zn_2SiO_4:Mn^{2+}$  particles. *Materials Chemistry and Physics*, 148(3), 1055–1063.
- Mohamed, N., Hassan, J., Matori, K. A., Wahab, Z. A., Ismail, Z. M. M., Baharuddin, N. F., and Rashid, S. S. A. (2017). Influence of Pr doping on the thermal, structural and optical properties of novel SLS-ZnO glasses for red phosphor. *Results in Physics*, 7, 1202-1206.

- Mohd Zaid, M. H., Amin Matori, K., Abdul Aziz, S. H., Kamari, H. M., Mat Yunus, W. M., Abdul Wahab, Z., and Samsudin, N. F. (2016). Fabrication and crystallization of ZnO-SLS glass derived willemite glass-ceramics as a potential material for optics applications. *Journal of Spectroscopy*, 1-7.
- Molla, A. R., Tarafder, A., Mukherjee, S., and Karmakar, B. (2011). Transparent Nd<sup>3+</sup>-doped bismuth titanate glass-ceramic nanocomposites: Fabrication and properties. *Optical Materials Express*, 4(4), 843-863.
- Mor, S., Ravindra, K., Dahiya, R. P., and Chandra, A. (2006). Leachate characterization and assessment of groundwater pollution near municipal solid waste landfill site. *Environmental monitoring and assessment*, 118(1), 435-456.
- Mumpton, F. A. (1956). Stability of minerals in the system ZnO-SiO<sub>2</sub>-H<sub>2</sub>O. *Economic Geology*, 51(5), 432-443.
- Noushad, M., Rahman, I. A., Husein, A., Mohamad, D., and Ismail, A. R. (2012). A Simple Method of Obtaining Spherical Nanosilica from Rice Husk. *International Journal on Advanced Science Information Technology*, 2(2), 28–30.
- Omar, N. A. S., Fen, Y. W., Matori, K. A., Zaid, M. H. M., Norhafizah, M. R., Nurzilla, M., and Zamratul, M. I. M. (2016). Synthesis and optical properties of europium doped zinc silicate prepared using low cost solid state reaction method. *Journal of Materials Science: Materials in Electronics*, 27(2), 1092-1099.
- Omri, K., El Ghoul, J., Alyamani, A., Barthou, C., and El Mir, L. (2013). Luminescence properties of green emission of SiO<sub>2</sub>/Zn<sub>2</sub>SiO<sub>4</sub>:Mn nanocomposite prepared by sol-gel method. *Physica E: Low-dimensional Systems and Nanostructures*, 53, 48-54.
- Omri, K., and El Mir, L. (2014). Effect of manganese concentration on photoluminescence properties of Zn<sub>2</sub>SiO<sub>4</sub>:Mn nano phosphor material. *Super lattices and Microstructures*, 70, 24–32.
- Omri, K., El Mir, L., Dahman, H., and Barthou, C. (2014). Synthesis and Luminescence Properties of Yellow-emitting SiO<sub>2</sub>/Zn<sub>2</sub>SiO<sub>4</sub>:Mn Nanocomposite. *Sensors and Transducers*, 27(5), 295.
- Ozel, E., Yurdakul, H., Turan, S., Ardit, M., Cruciani, G., and Dondi, M. (2010). Co-doped willemite ceramic pigments: Technological behaviour, crystal structure and optical properties. *Journal of the European Ceramic Society*, 30(16), 3319–3329.
- Pai, P. G., Chao, S. S., Takagi, Y., and Lucovsky, G. (1986). Infrared spectroscopic study of SiO<sub>x</sub> films produced by plasma enhanced chemical vapor deposition. *Journal of Vacuum Science and Technology A: Vacuum, Surfaces, and Films*, 4(3), 689-694.
- Palmer, D. A., and Anovitz, L. M. (2009). Solubility of zinc silicate and zinc ferrite in aqueous solution to high temperatures. *Journal of Solution Chemistry*, 38(7), 869–892.

- Parhi, P., and Manivannan, V. (2009). Novel microwave initiated synthesis of  $Zn_2SiO_4$  and  $MCrO_4$  ( $M=Ca, Sr, Ba, Pb$ ). *Journal of Alloys and Compounds*, 469(1-2), 558–564.
- Pascuta, P., Rada, S., Borodi, G., Bosca, M., Pop, L., and Culea, E. (2009). Influence of europium ions on structure and crystallization properties of bismuth-alumino-borate glasses and glass ceramics. *Journal of Molecular Structure*, 924, 214-220
- Pattengil, M., and Shutt, T. (1973). Use of ground glass as a pozzolan; *Albuquerque symposiums on Utilization of Waste Glass in Secondary Products*. Albuquerque, New Mexico, USA.
- Patra, a., Baker, G. A., and Baker, S. N. (2004). Synthesis and luminescence study of  $Eu^{3+}$  in  $Zn_2SiO_4$  nanocrystals. *Optical Materials*, 27(1), 15–20.
- Patra, a., Baker, G. A., and Baker, S. N. (2005). Effects of dopant concentration and annealing temperature on the phosphorescence from  $Zn_2SiO_4:Mn^{2+}$  nanocrystals. *Journal of Luminescence*, 111(1-2), 105–111.
- Petrescu, S., Constantinescu, M., Anghel, E. M., Atkinson, I., Olteanu, M., and Zaharescu, M. (2012). Structural and physico-chemical characterization of some soda lime zinc alumino-silicate glasses. *Journal of Non-Crystalline Solids*, 358(23), 3280–3288.
- Pfeiffer, H. G., and Fonda, G. R. (1952). The zinc silicate phosphors fluorescing in the yellow and red. *Journal of the Electrochemical Society*, 99(4), 140-143.
- Philipp, H. R., and Ehrenreich, H. (1963). Optical properties of semiconductors. *Physical Review*, 129(4), 1550–1560.
- Phillips, J.C. and Cahn, D.S. (1973). Refuse Glass Aggregate in Portland Cement. Proc. 3rd Mineral Waste Utilisation Symposium, pp. 385-390.
- Podbršček, P., Dražić, G., Anžlovar, A., and Orel, Z. C. (2011). The preparation of zinc silicate/ $ZnO$  particles and their use as an efficient UV absorber. *Materials Research Bulletin*, 46(11), 2105-2111.
- Qazvini, S. S. A., Hamnabard, Z., Khalkhali, Z., Baghshahi, S., and Maghsoudipour, A. (2012). Photoluminescence and microstructural properties of  $SiO_2-ZnO-B_2O_3$  system containing  $TiO_2$  and  $V_2O_5$ . *Ceramics International*, 38(2), 1663–1670.
- Qu, J., Cao, C. Y., Hong, Y. L., Chen, C. Q., Zhu, P. P., Song, W. G., and Wu, Z. Y. (2012). New hierarchical zinc silicate nanostructures and their application in lead ion adsorption. *Journal of Materials Chemistry*, 22(8), 3562-3567.
- Rahman, I. A., and Riley, F. L. (1989). The control of morphology in silicon nitride powder prepared from rice husk. *Journal of the European Ceramic Society*, 5(1), 11–22.
- Rani, S., Sanghi, S., Ahlawat, N., and Agarwal, A. (2015). Influence of  $Bi_2O_3$  on physical, electrical and thermal properties of  $Li_2O \cdot ZnO \cdot Bi_2O_3 \cdot SiO_2$  glasses. *Journal of Alloys and Compounds*, 619, 659-666.
- Rao, P. S., Rajyasree, C., Babu, A. R., Teja, P. V., and Rao, D. K. (2011). Effect of  $Bi_2O_3$  proportion on physical, structural and electrical properties of zinc

- bismuth phosphate glasses. *Journal of Non-Crystalline Solids*, 357(21), 3585-3591.
- Rashid, S. S. A., Ab Aziz, S. H., Matori, K. A., Zaid, M. H. M., and Mohamed, N. (2017). Comprehensive study on effect of sintering temperature on the physical, structural and optical properties of  $\text{Er}^{3+}$  doped  $\text{ZnO-GSLs}$  glasses. *Results in Physics*.
- Reddy, C. P., Naresh, V., and Babu, B. C. (2014). Photoluminescence and Energy Transfer Process in  $\text{Bi}^{3+}/\text{Sm}^{3+}$  Co-Doped Phosphate Zinc Lithium Glasses. *Advances in Materials Physics and Chemistry*, 4(9), 165–171.
- Ringwood, A. E., and Major, A. (1967). High pressure transformations in zinc germanates and silicates. *Nature*, 215(5108), 1367-1368.
- Rooksby, H. P., and McKeag, A. H. (1941). The yellow fluorescent form of zinc silicate. *Transactions of the Faraday Society*, 37, 308-311.
- Rosenthal, A. B., and Garofalini, S. H. (1986). Molecular dynamics simulation of amorphous zinc silicate. *Journal of non-crystalline solids*, 87(1-2), 254-262.
- Ruiterkamp, R., Halasinski, T., Salama, F., Foing, B. H., Allamandola, L. J., Schmidt, W., and Ehrenfreund, P. (2002). Spectroscopy of large PAHs-Laboratory studies and comparison to the diffuse interstellar bands. *Astronomy and Astrophysics*, 390(3), 1153-1170.
- Saarela, J. (2003). Pilot investigations of surface parts of three closed landfills and factors affecting them. *Environmental monitoring and assessment*, 84(1), 183-192.
- Sahu, A. K., Kumar, D., Parkash, O., Thakur, O. P., and Prakash, C. (2004). Effect of  $\text{K}_2\text{O}/\text{BaO}$  ratio on crystallization, microstructure and dielectric properties of strontium titanate borosilicate glass ceramics. *Ceramics international*, 30(3), 477-483.
- Sahu, I. P., Bisen, D. P., and Brahme, N. (2015). Luminescence properties of green-emitting  $\text{Ca}_2\text{MgSi}_2\text{O}_7:\text{Eu}^{2+}$  phosphor by a solid-state reaction method. *Luminescence*, 30(7), 1125-1132.
- Sahu, I. P., Bisen, D. P., and Brahme, N. (2015). Structural characterization and optical properties of  $\text{Ca}_2\text{MgSi}_2\text{O}_7:\text{Eu}^{2+}, \text{Dy}^{3+}$  phosphor by solid-state reaction method. *Luminescence*, 30(5), 526–532.
- Sahu, I. P., Bisen, D. P., Brahme, N., and Tamrakar, R. K. (2015). Structural characterization and optical properties of dysprosium doped strontium calcium magnesium di-silicate phosphor by solid state reaction method Ishwar. *Displays*, 8(1), 104–109.
- Saikia, B. J., Parthasarathy, G., and Sarmah, N. C. (2008). Fourier transform infrared spectroscopic estimation of crystallinity in  $\text{SiO}_2$  based rocks. *Bulletin of Materials Science*, 31(5), 775-779.
- Samsudin, N. F., Matori, K. A., Liew, J. Y. C., Wing Fen, Y., Mohd Zaid, M. H., and Nadakkavil Alassan, Z. (2015). Investigation on Structural and Optical Properties of Willemite Doped  $\text{Mn}^{2+}$ . *Journal of Spectroscopy*, 2015(2015), 1-7.

- Samtur, H. R. (1974). Glass recycling and reuse (No. PB-239674). Wisconsin Univ., Madison (USA). Inst. for Environmental Studies.
- Sánchez-Royo, J. F., Muñoz-Matutano, G., Brotons-Gisbert, M., Martínez-Pastor, J. P., Segura, A., Cantarero, A., Gerardot, B. D. (2014). Electronic structure, optical properties, and lattice dynamics in atomically thin indium selenide flakes. *Nano Research*, 7(10), 1556–1568.
- Schleede, A., and Gruhl, A. (1923). Über Röntgenographische Beobachtungen an lumineszenzfähigem Zinksilikat. *Berichte der Bunsengesellschaft für physikalische Chemie*, 29(17-18), 411-412.
- Scherer, G. W. (1997). Sintering of sol-gel films. *Journal of Sol-Gel Science and Technology*, 8(1-3), 353-363.
- Schneider, J., Boni, M., Laukamp, C., Bechstädt, T., and Petzel, V. (2008). Willemite ( $Zn_2SiO_4$ ) as a possible Rb-Srgeochronometer for dating nonsulfide Zn-Pb mineralization: Examples from the OtaviMountainland (Namibia). *Ore Geology Reviews*, 33(2), 152–167.
- See, A., Hassan, J., Hashim, M., and Wahab, Z. A. (2014). Thermal diffusivity of kaolinite-mullite ceramic matrix composite with silicon nitride nanoparticle filler. *Thermochimica Acta*, 593, 76-81.
- Selomulya, R., Ski, S., Pita, K., Kam, C. ., Zhang, Q. ., and Buddhudu, S. (2003). Luminescence properties of  $Zn_2SiO_4:Mn^{2+}$  thin-films by a sol-gel process. *Materials Science and Engineering: B*, 100(2), 136–141.
- Selvi, S., Marimuthu, K., and Muralidharan, G. (2015). Structural and luminescence behavior of  $Sm^{3+}$  ions doped lead boro-telluro-phosphate glasses. *Journal of Luminescence*, 159, 207–218.
- Sembiring, E., and Nitivattananon, V. (2010). Sustainable solid waste management toward an inclusive society: Integration of the informal sector. *Resources, Conservation and Recycling*, 54(11), 802-809.
- Shalaev, E. Y., and Franks, F. (1995). Structural glass transitions and thermophysical processes in amorphous carbohydrates and their supersaturated solutions. *J. Chem. Soc., Faraday Trans.*, 91(10), 1511-1517.
- Sharma, P., and Bhatti, H. S. (2009). Laser induced down conversion optical characterizations of synthesized  $Zn_{2-x}Mn_xSiO_4$  ( $0.5 \leq x \leq 5$  mol%) nanophosphors. *Journal of Alloys and Compounds*, 473(1-2), 483–489.
- Shayan, A., and Xu, A. (1999). Utilization of glass as a pozzolanic material in concrete. ARRB TR Internal Report RC91132.
- Shayan, A., and Xu, A. (2004). Value-added utilisation of waste glass in concrete. *Cement and concrete research*, 34(1), 81-89.
- Shelby, J. E. (1981). Effect of crystal content on the properties of willemite glass-ceramics. *Journal of Non-Crystalline Solids*, 43(2), 255-265.
- Sreedhar, V. B., Basavapoornima, C., and Jayasankar, C. K. (2014). Spectroscopic and fluorescence properties of  $Sm^{3+}$ -doped zincfluorophosphate glasses. *Journal of Rare Earths*, 32(10), 918–926.

- Srivastava, A. M. (2002). On the luminescence of  $\text{Bi}^{3+}$  in the pyrochlore  $\text{Y}_2\text{Sn}_2\text{O}_7$ . *Materials Research Bulletin*, 37, 745–751.
- Sułowska, J., Wacławska, I., and Szumera, M. (2016). Comparative study of zinc addition effect on thermal properties of silicate and phosphate glasses. *Journal of Thermal Analysis and Calorimetry*, 123(2), 1091–1098.
- Sweet, J. R., Hess, J. W., and White, W. B. (2010). Candaluminescence of cave gypsum. *International Journal of Speleology*, 39(1), 25–28.
- Syamimi, N. F., Amin Matori, K., Lim, W. F., Abdul Aziz, S., and MohdZaid, M. H. (2014). Effect of sintering temperature on structural and morphological properties of europium (III) oxide doped willemite. *Journal of Spectroscopy*, 2014(2014), 1–9.
- Syono, Y., Akimoto, S. I., and Matsui, Y. (1971). High pressure transformations in zinc silicates. *Journal of Solid State Chemistry*, 3(3), 369–380.
- Taghavinia, N., Lerondel, G., Makino, H., Yamamoto, A., Yao, T., Kawazoe, Y., and Goto, T. (2001). Nanocrystalline  $\text{Zn}_2\text{SiO}_4:\text{Mn}^{2+}$  grown in oxidized porous silicon. *Nanotechnology*, 12(4), 547.
- Takesue, M., Hayashi, H., and Smith, R. L. (2009). Thermal and chemical methods for producing zinc silicate (willemite): A review. *Progress in Crystal Growth and Characterization of Materials*, 55(3), 98–124.
- Tanabe, S., Hayashi, H., Hanada, T., and Onodera, N. (2002). Fluorescence properties of  $\text{Er}^{3+}$  ions in glass ceramics containing  $\text{LaF}_3$  nanocrystals. *Optical Materials*, 19(3), 343–349.
- Tang, X., Choo, E. S. G., Li, L., Ding, J., and Xue, J. (2010). Synthesis of  $\text{ZnO}$  nanoparticles with tunable emission colors and their cell labeling applications. *Chemistry of Materials*, 22(11), 3383–3388.
- Tani, T., Watanabe, N., and Takatori, K. (2003). Emulsion combustion and flame spray synthesis of zinc oxide/silica particles. *Journal of Nanoparticle Research*, 5(1), 39–46.
- Tarafder, A., Molla, A. R., Dey, C., and Karmakar, B. (2013). Thermal, Structural, and Enhanced Photoluminescence Properties of  $\text{Eu}^{3+}$  doped Transparent Willemite Glass–Ceramic Nano composites. *Journal of the American Ceramic Society*, 96(8), 2424–2431.
- Tarafder, A., Molla, A. R., Mukhopadhyay, S., and Karmakar, B. (2014). Fabrication and enhanced photoluminescence properties of  $\text{Sm}^{3+}$  doped  $\text{ZnO}-\text{Al}_2\text{O}_3-\text{B}_2\text{O}_3-\text{SiO}_2$  glass derived willemite glass-ceramic nanocomposites. *Optical Materials*, 36(9), 1463–1470.
- Taylor, H. F. W. (1962). The dehydration of hemimorphite. *American Mineral*, 47, 932–944.
- Thomas, S., and Sebastian, M. T. (2009). Microwave Dielectric Properties of  $\text{SrRE}_4\text{Si}_3\text{O}_{13}$  ( $\text{RE}=\text{La}, \text{Pr}, \text{Nd}, \text{Sm}, \text{Eu}, \text{Gd}, \text{Tb}, \text{Dy}, \text{Er}, \text{Tm}, \text{Yb}$ , and  $\text{Y}$ ) Ceramics. *Journal of the American Ceramic Society*, 92(12), 2975–2981.
- Thomas, S., Rasool, S. N., Rathaiah, M., Venkatramu, V., Joseph, C., and Unnikrishnan, N. V. (2013). Spectroscopic and dielectric studies of  $\text{Sm}^{3+}$  ions

- in lithium zinc borate glasses. *Journal of Non-Crystalline Solids*, 376, 106–116.
- Thuadaij, N., and Nuntiya, A. (2008). Synthesis and characterization of nanosilica from rice husk ash prepared by precipitation method. *J. Nat. Sci. Special Issue on Nanotechnology*, 7(1), 59-65.
- Todea, M., Turcu, R. V. F., Vasilescu, M., Trandafir, D. L., and Simon, S. (2016). Structural characterization of heavy metal  $\text{SiO}_2\text{--Bi}_2\text{O}_3$  glasses and glass-ceramics. *Journal of Non-Crystalline Solids*, 432, 271-276.
- Topcu, I. B., and Canbaz, M. (2004). Properties of concrete containing waste glass. *Cement and Concrete Research*, 34(2), 267-274.
- Torres, F. J., Rodríguez-Mendoza, U. R., Lavín, V., de Sola, E. R., and Alarcón, J. (2007). Evolution of the structural and optical properties from cobalt cordierite glass to glass-ceramic based on spinel crystalline phase materials. *Journal of Non-Crystalline Solids*, 353(44-46), 4093–4101.
- Tummala, R. R. (1991). Ceramic and Glass-Ceramic Packaging in the 1990s. *Journal of the American Ceramic Society*, 74(5), 895-908.
- Uegaito, K., Hosokawa, S., and Inoue, M. (2011). Effect of heat treatments on the luminescence properties of  $\text{Zn}_2\text{SiO}_4:\text{Mn}^{2+}$  phosphors prepared by glycothermal methods. *Journal of Luminescence*, 132(1), 64–70.
- Ura, C., Cedex, G. Y., Soleil, S., Yvette, G., Uesms, I., and Saclay, C. E. A. (2010). Silicon and Titanium Oxide Nanoparticles by Laser Pyrolysis : Elaboration on Nanostructured Materials for Optical and Energy Applications, 7, 552–558.
- Van, L., Melis, G. P., Brandsma, T. C. 2000.Thermal Conductivity of Glass: Measurement by the melt contact.Glass Physics and Chemistry.Vol 26:6
- Vanderbilt, D. (1990). Soft self-consistent pseudopotentials in a generalized eigenvalue formalism. *Physical Review B: Rapid Communications*, 41(11), 7892–7895.
- Venkatramu, V., Babu, P., Jayasankar, C. K., Tröster, T., Sievers, W., and Wortmann, G. (2007).Optical spectroscopy of  $\text{Sm}^{3+}$  ions in phosphate and fluorophosphate glasses. *Optical Materials*, 29(11), 1429–1439.
- Wagh, A., Raviprakash, Y., Ajithkumar, M. P., Upadhyaya, V., and Kamath, S. D. (2015). Effect of  $\text{Sm}_2\text{O}_3$  on structural and thermal properties of zinc fluoroborate glasses. *Transactions of Nonferrous Metals Society of China*, 25(4), 1185–1193.
- Wan, J., Chen, X., Wang, Z., Mu, L., and Qian, Y. (2005). One-dimensional rice-like Mn-doped  $\text{Zn}_2\text{SiO}_4$ : Preparation, characterization, luminescent properties and its stability. *Journal of crystal growth*, 280(1), 239-243.
- Wan, J., Wang, Z., Chen, X., Mu, L., Yu, W., and Qian, Y. (2006). Controlled synthesis and relationship between luminescent properties and shape/crystal structure of  $\text{Zn}_2\text{SiO}_4:\text{Mn}^{2+}$  phosphor. *Journal of luminescence*, 121(1), 32-38.
- Wang, L., Liu, X., Hou, Z., Li, C., Yang, P., Cheng, Z., Lian, H., and Lin, J. (2008). Electrospinning synthesis and luminescence properties of one-dimensional

- Zn<sub>2</sub>SiO<sub>4</sub>:Mn<sup>2+</sup> microfibers and microbelts. *The Journal of Physical Chemistry C*, 112(48), 18882-18888.
- Weidenkaff, A., Reller, A., and Steinfeld, A. (2001). Solar Production of Zinc from the Zinc Silicate Ore Willemite. *Journal of Solar Energy Engineering*, 123(2), 98-101.
- Williamson, J., and Glasser, F. P. (1964). Crystallisation of zinc silicate liquids and glasses. *Phys. Chem. Glasses*, 5(1), 52-59.
- Wu, P., Eriksson, G., and Pelton, A. D. (1993). Optimization of the thermodynamic properties and phase diagrams of the Na<sub>2</sub>O-SiO<sub>2</sub> and K<sub>2</sub>O-SiO<sub>2</sub> systems. *Journal of the American Ceramic Society*, 76(8), 2059-2064.
- Wu, S., Song, K., Liu, P., Lin, H., Zhang, F., Zheng, P., and Qin, H. (2015). Effect of TiO<sub>2</sub> Doping on the Structure and Microwave Dielectric Properties of Cordierite Ceramics. *Journal of the American Ceramic Society*, 98(6), 1842-1847.
- Yakuphanoglu, F., Ilican, S., Caglar, M., and Caglar, Y. (2007). The determination of the optical band and optical constants of non-crystalline and crystalline ZnO thin films deposited by spray pyrolysis. *Journal of Optoelectronics and Advanced Materials*, 9(7), 2180-2185.
- Yamashita, H., Honda, M., Harada, M., Ichihashi, Y., Anpo, M., Hirao, T. and Iwamoto, N. (1998). Preparation of titanium oxide photocatalysts anchored on porous silica glass by a metal ion-implantation method and their photocatalytic reactivities for the degradation of 2-propanol diluted in water. *The Journal of Physical Chemistry B*, 102(52), 10707-10711.
- Yang, P., Lü, M. K., Song, C. F., Xu, D., Yuan, D. R., Xia, G. M., Shen, J. X. (2003). Photoluminescence of Bi<sup>3+</sup> ions in sol-gel derived Zn<sub>2</sub>SiO<sub>4</sub>. *Materials Research Bulletin*, 38(5), 757-763.
- Yang, Y., Zhuang, Y., He, Y., Bai, B., and Wang, X. (2010). Fine tuning of the dimensionality of zinc silicate nanostructures and their application as highly efficient absorbents for toxic metal ions. *Nano Research*, 3(8), 581-593.
- Yekta, B. E., Alizadeh, P., and Rezazadeh, L. (2007). Synthesis of glass-ceramic glazes in the ZnO-Al<sub>2</sub>O<sub>3</sub>-SiO<sub>2</sub>-ZrO<sub>2</sub> system. *Journal of the European Ceramic Society*, 27(5), 2311-2315.
- Youngblood, G.E., Rice, R.W. and Ingel, R.P. 1988. Thermal diffusivity of partially and fully stabilized (yttria) zirconia single crystals. *J. Am. Ceram. Soc.* 71(4):255-260.
- Yurieva, T. M., Kustova, G. N., Minyukova, T. P., Poels, E. K., Bliek, A., Demeshkina, M. P., and Zaikovskii, V. I. (2001). Non-hydrothermal synthesis of copper-, zinc-and copper-zinc hydrosilicates. *Materials Research Innovations*, 5(1), 3-11.
- Zaid, M. H. M., Matori, K. A, Wah, L. C., Sidek, H. A.A., Halimah, M. K., Wahab, Z. A, and Azmi, B. Z. (2011). Elastic moduli prediction and correlation in soda lime silicate glasses containing ZnO. *International Journal of the Physical Sciences*, 6(6), 1404-1410.

- Zaid, M. H. M., Matori, K. A., Aziz, S. H. A., Kamari, H. M., Wahab, Z. A., Fen, Y. W., and Alibe, I. M. (2016). Synthesis and characterization of low cost willemite based glass-ceramic for opto-electronic applications. *Journal of Materials Science: Materials in Electronics*, 27(11), 11158-11167.
- Zaid, M. H. M., Matori, K. A., Wah, L. C., Sidek, H. A. A., Halimah, M. K., Wahab, Z. A., and Azmi, B. Z. (2011). Elastic moduli prediction and correlation in soda lime silicate glasses containing ZnO. *International Journal of Physical Sciences*, 6(6), 1404-1410.
- Zainul, H., Talib, Z. A., Zaidan, A. W., Zainal, A. S., and Mansor, H. (2005). Ultrasonic study and Physical Properties of Borotellurite Glasses. *American Journal of Applied Sciences*, 2(11), 1541–1546.
- Zanotto, E. D. (2010). A bright future for glass-ceramics. *American Ceramic Society Bulletin*, 89(8), 19–27.
- Zhang, H. X., Buddhudu, S., Kam, C. H., Zhou, Y., Lam, Y. L., Wong, K. S., and Que, W. X. (2001). Luminescence of Eu<sup>3+</sup> and Tb<sup>3+</sup> doped Zn<sub>2</sub>SiO<sub>4</sub> nanometer powder phosphors. *Materials chemistry and physics*, 68(1), 31-35.
- Zhang, H. X., Kam, C. H., Zhou, Y., Han, X. Q., Buddhudu, S., Lam, Y. L., and Chan, C. Y. (2000). Deposition and photoluminescence of sol±gel derived Tb<sup>+3</sup>: Zn<sub>2</sub>SiO<sub>4</sub> on SiO<sub>2</sub>/Si. *Thin Solid Films*, 370, 50–53.
- Zhang, M., Lin, K., and Chang, J. (2012). Preparation and characterization of Sr-hardystonite (Sr<sub>2</sub>ZnSi<sub>2</sub>O<sub>7</sub>) for bone repair applications. *Materials Science and Engineering: C*, 32(2), 184–188.
- Zhang, Q. Y., Pita, K., Ye, W., and Que, W. X. (2002). Influence of annealing atmosphere and temperature on photoluminescence of Tb<sup>3+</sup> or Eu<sup>3+</sup> activated zinc silicate thin film phosphors via sol-gel method. *Chemical Physics Letters*, 351(3–4), 163–170.
- Zhao, W., Li, S., Zhang, S., Liu, X., and Hou, J. (2017). Ternary polymer solar cells based on two acceptors and one donor for achieving 12.2% efficiency. *Advanced Materials*, 29(2).
- Zhong, L. S., Hu, J. S., Liang, H. P., Cao, A. M., Song, W. G., and Wan, L. J. (2006). Self-Assembled 3D flowerlike iron oxide nanostructures and their application in water treatment. *Advanced Materials*, 18(18), 2426-2431.