

## **UNIVERSITI PUTRA MALAYSIA**

# ELASTIC PROPERTIES AND IONIC CONDUCTIVITY OF LITHIUM PHOSPHATE AND LITHIUM BORATE GLASSES

**LOW YEE SAN** 

FSAS 1999 9



# ELASTIC PROPERTIES AND IONIC CONDUCTIVITY OF LITHIUM PHOSPHATE AND LITHIUM BORATE GLASSES

**LOW YEE SAN** 

MASTER OF SCIENCE UNIVERSITI PUTRA MALAYSIA

1999



# ELASTIC PROPERTIES AND IONIC CONDUCTIVITY OF LITHIUM PHOSPHATE AND LITHIUM BORATE GLASSES

Ву

LOW YEE SAN

Thesis Submitted in Fulfilment of the Requirements for the Degree of Master of Science in the Faculty of Science and Environmental Studies,
Universiti Putra Malaysia

June 1999



### **ACKNOWLEDGEMENT**

First and foremost, I wish to thank my supervisor, Associate Professor Dr. Sidek Hj. Abd. Aziz for his supportive and friendly supervision. It is also my great privilege to have Associate Professor Dr. Chow Sai Pew as my co-supervisor. He has been kind and helpful to me over the past two years and I can never thank him enough for all his help and guidance.

I would like to extend my appreciation to Associate Professor Dr. Hj. Wan Mohd Daud Wan Yusoff and Dr. Mansor Hashim for their kindness in allowing me to explore the possibility of using their equipment for the ionic conductivity measurement. I would also wish to thank Associate Professor Dr. Zainal Abidin Talib for the beneficial discussions.

I would like to extend my gratitude to the government for granting the PASCA scholarship to me to support my daily expenditure. Finally, thanks to my heavenly Father for His grace and mercy throughout this two years.



## TABLE OF CONTENT

		Page
	IOWLEDGEMENTS	11
	OF TABLES	vii
	OF FIGURES	1X
	OF PLATES	xvi
	OF ABBREVIATIONS	xvii
	RACT	xviii
ABST	RAK	XX
СНАР	PTER	
I	RESEARCH OVERVIEW	1
	Introduction	1
	Literature Review	2
	Phosphate Glasses	2
	Borate Glasses	3
	Objectives	5
	Chapter Organization	5
II	THEORY OF GLASS	7
	Introduction	7
	Fusion and Crystallization	9
	Volume- Temperature Relationship	11
	Glass Formation	13
	Goldschmidt's Radius Ratio Criteria	14
	Zachariasen's Random Network Theory	15
	Dietzel's Field Strength Criterion	20
	Sun's Single Bond Strength Criterion	21
	Smekal's Mixed Bonding Rule	23
	Kinetic Theory	25
	Phosphate Glasses	27
	Borate Glasses	30
	Boron Anomaly	32
	Structure of Alkali Borate Glasses	34
	Chapter Summary	36



## III THEORY OF LINEAR

	Introduction	38
	Stress	39
	Strain	41
	The Perfectly Elastic Body	43
	The Isotropic Body	44
	Propagation of Elastic Wave in an Isotropic Solid	46
	Equation of Motion	48
	Elastic Wave Velocities in Isotropic Medium	49
	Ultrasonic Velocities and Elastic Modulus	52
	Bulk Modulus	52
	Young's Modulus	53
	Poisson's Ratio	53
	Debye Temperature	54
	Chapter	54
V	THEORY OF IONIC	56
	Introduction	
	Mechanism for Ionic Transport	51
	Anderson-Stuart	
	Weak-Electrolyte	
	Defect-Based	
	Frequency-Dependent	
	Composition-Dependent	
	Chapter Summary	66
V	METHODOLOGY	68
	Introduction	68
	Glass Preparation	68
	Melting	69
	Casting of the Melt	70
	Glass Cutting and Polishing	74
	Measurement of Density	76
	Ultrasonic Measurement	77
	MBS-8000	
	Ultrasonic Transducer	83
	Measurement Procedures	84



	Ionic Conductivity Measurement	85
	Measurement Instruments	85
	Measurement Conditions	87
	Measurement Connections	88
	Experimental Errors	89
	Chapter Summary	90
VI	ELASTIC PROPERTIES OF PHOSPHATE GLASSES	91
	Introduction	91
	Results and Discussion	91
	Density	94
	Molar Volume	95
	Ultrasonic Velocities	97
	Glass Structure	98
	Elastic Moduli	102
	Debye Temperature	107
	Summary	109
VII	ELASTIC PROPERTIES OF BORATE GLASSES	110
	Introduction	110
	Results and Discussion	111
	Structure of Borate Glasses	111
	Structure of Lithium Borate Glasses	114
	Density	117
	Molar Volume	119
	Discrepancy of Density and Molar Volume	121
	Ultrasonic Velocities	122
	Elastic Moduli	125
	Debye Temperature	130
	Summary	131
VIII	IONIC PROPERTIES OF PHOSPHATE GLASSES	132
	Introduction	132
	Results and Discussion	133
	Frequency-Dependent Ionic Conductivity	133
	Temperature-Dependent Ionic Conductivity	143



	Summary	
IX	IONIC PROPERTIES OF BORATE GLASSES	151
	Introduction	151
	Results and Discussion	
	Frequency-Dependent Ionic Conductivity	152
	Temperature-Dependent Ionic Conductivity	158
	Composition-Dependent Ionic Conductivity	160
	Summary	165
X	GENERAL DISCUSSION AND CONCLUSION	166
	Introduction	166
	Discussion	166
	Conclusion	172
	Suggestion	174
BIBLIC	OGRAPHY	176
APPEN	NDIX	
A	Vector Operators	182
В	Plates of Apparatus	184
C	Chemical Proportion of Glass Samples	189
D	Results of Ionic Conductivity Measurement	
VITA		204



## LIST OF TABLES

Table		Page
2.1	Radius for Typical Glass Formers	15
2.2	Classification of Cations According to Their Field Strength	22
2.3	Single Bond Strength for Oxides	24
6.1	Density, Molar Volume, Longitudinal and Shear Velocities, Elastic Moduli and Debye Temperature of $(Li_2O)_x(P_2O_5)_{1-x}$ Glasses at Room Temperature	92
6.2	Density, Molar Volume, Longitudinal and Shear Velocities, Elastic Moduli and Debye Temperature of (LiCl) <sub>y</sub> [(Li <sub>2</sub> O) <sub>0.4</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0.6</sub> ] <sub>1-y</sub> Glasses at Room Temperature	93
7.1	Density, Molar Volume, Longitudinal and Shear Velocities, Elastic Moduli and Debye Temperature of $(Li_2O)_x(B_2O_3)_{1-x}$ Glasses at Room Temperature	112
7.2	Density, Molar Volume, Longitudinal and Shear Velocities, Elastic Moduli and Debye Temperature of $(\text{LiCl})_y[(\text{Li}_2O)_{0.25}(P_2O_5)_{0.75}]_{1-y}$ Glasses at Room Temperature	113
7.3	Structure Groups Present at Compositions of xLi <sub>2</sub> O(100-x)B <sub>2</sub> O <sub>3</sub> Glasses	116
Result	s of Ionic Conductivity Measurement for:	
D1	$(\text{Li}_2\text{O})_x(\text{P}_2\text{O}_5)_{1-x}$ Glass With $x = 0.10$	191
D2	$(Li_2O)_x(P_2O_5)_{1-x}$ Glass With $x = 0.15$	192
D3	$(Li_2O)_x(P_2O_5)_{1-x}$ Glass With $x = 0.20$	192
D4	$(\text{Li}_2\text{O})_x(\text{P}_2\text{O}_5)_{1-x}$ Glass With $x = 0.25$	193
D5	$(\text{Li}_2\text{O})_x(\text{P}_2\text{O}_5)_{\text{Li}}$ Glass With $x = 0.30$	193



D6	$(L_{12}O)_x(P_2O_5)_{1-x}$ Glass With $x = 0.35$	194
D7	$(Li_2O)_x(P_2O_5)_{1-x}$ Glass With $x = 0.40$	194
D8	$(Li_2O)_x(P_2O_5)_{1-x}$ Glass With $x = 0.45$	195
D9	$(Li_2O)_x(P_2O_5)_{1-x}$ Glass With $x = 0.50$	195
D10	$(LiCl)_y[(Li_2O)_0 _4(P_2O_5)_0 _6]_{1-y}$ Glass With $y = 0 10$	196
D11	$(LiCl)_y[(Li_2O)_0_4(P_2O_5)_0_6]_{1-y}$ Glass With $y = 0.20$	196
D12	$(LiCl)_y[(Li_2O)_{0.4}(P_2O_5)_{0.6}]_{1-y}$ Glass With $y = 0.30$	197
D13	$(LiCl)_y[(Li_2O)_0_4(P_2O_5)_0_6]_{1-y}$ Glass With $y = 0.40$	197
D14	$(LiCl)_y[(Li_2O)_{0.4}(P_2O_5)_{0.6}]_{1-y}$ Glass With $y = 0.50$	198
D15	$(Li_2O)_x(B_2O_3)_{1-x}$ Glass With $x = 0.15$	198
D16	$(Li_2O)_x(B_2O_3)_{1-x}$ Glass With $x = 0.20$	199
D17	$(Li_2O)_x(B_2O_3)_{1-x}$ Glass With $x = 0.25$	199
D18	$(Li_2O)_x(B_2O_3)_{1-x}$ Glass With $x = 0 30$	200
D19	$(Li_2O)_x(B_2O_3)_{1-x}$ Glass With $x = 0.35$	200
D20	$(Li_2O)_x(B_2O_3)_{1-x}$ Glass With $x = 0.40$	201
D21	$(LiCl)_y[(Li_2O)_{0.25}(B_2O_3)_{0.75}]_{1-y}$ Glass With $y = 0.05$	201
D22	$(LiCl)_y[(Li_2O)_{0.25}(B_2O_3)_{0.75}]_{1-y}$ Glass With $y = 0.10$	202
D23	$(LiCl)_y[(L_{12}O)_{0.25}(B_2O_3)_{0.75}]_{1-y}$ Glass With $y = 0.15$	202
D24	$(LiCl)_y[(Li_2O)_{0.25}(B_2O_3)_{0.75}]_{1-y}$ Glass With $y = 0.20$	203
D25	(LiCl) $[(Li_2O)_{0.05}(B_2O_2)_{0.05}]$ . Glass With $y=0.25$	203



## LIST OF FIGURES

Figure		Page	
2.1	The Volume-temperature diagram for a glass-forming liquid	11	
2.2	Atomic structural representation of (a) A <sub>2</sub> O <sub>3</sub> crystal (b) A <sub>2</sub> O <sub>3</sub> glass	17	
2.3	Structural representation of hypothetical crystalline compound AO	17	
2.4	Schematic illustration of the effect of the addition of alkali (e.g. Na <sub>2</sub> O) to silica	19	
2.5	T-T-T ("nose curve"). Plot of temperature vs time required to obtain 10-6 of the volume crystallized	26	
2.6	Structure of PO <sub>4</sub>	29	
2.7	Structure of P <sub>4</sub> O <sub>10</sub> molecules	29	
2.8	P <sup>5+</sup> and Al <sup>3+</sup> compensated in substituting Si <sup>4+</sup>	30	
2.9	The creation of non-bridging oxygen by adding alkali oxides	31	
2.10	Creation of BO <sub>3</sub> group to BO <sub>4</sub> group by adding alkali oxide	31	
2.11	Change of the coefficient of the linear thermal expansion as a function of composition in binary lithium, sodium, and potassium borate glasses	33	
2.12	The fraction N <sub>4</sub> of four coordinated boron atoms in alkali borate glasses	33	
2 13	Various structural groups present in several horate compounds	35	



3.1	The stress components acting on each face of an elementary parallelepiped when a uniform force acts on the body	40
3.2	An element PQ of an elastic body becomes the element P'Q' after deformation	42
3.3	Elastic waves in an unbounded isotropic medium (a) Longitudinal wave (b) Shear wave	4'
4.1	Configuration-coordinate diagram (a) Anderson-Stuart Model (b) Weak-electrolyte Model	59
4.2	Variation of conductivity with frequency over a wide range of temperature for Na <sub>2</sub> O.SiO <sub>2</sub> glass	6:
5.1	The steel cylindrical split moulds	7
5.2	Process involved in glass preparation	7
5.3	Glass forming range of LiCl-Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> system	7
5.4	Glass forming range of LiCl-Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> system	7
5.5	MBS-8000 system configuration block diagram	7
5.6	Time relationship of typical signals	8
5.7	MBS-8000 measurement system block diagram	8
5.8	Mechanical displacement of piezoelectric quartz plate operating in (a) the thickness shear mode (b) the thickness longitudinal mode	8
5.9	Four-terminal pair measurement principle	8
5.10	The configuration for clamping of glass samples	8
6.1	Composition dependence of density for Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> and LiCl-Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> glasses	0



6.2	Composition dependence of Molar Volume for Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> and LiCl-Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> glasses	96
6.3	Composition dependence of longitudinal and shear wave velocities for Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> and LiCl-Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> glasses	97
6.4	Proposed crosslinking effect of Li <sup>+</sup>	99
6.5	Structural model for LiCl-Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> glasses	101
6.6	Composition dependence of Young's modulus for Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> and LiCl-Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> glasses	103
6.7	Composition dependence of bulk modulus for Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> and LiCl-Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> glasses	104
6.8	Composition dependence of Poisson's ratio for Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> and LiCl-Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> glasses	107
6.9	Composition dependence of Debye temperature for Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> and LiCl-Li <sub>2</sub> O-P <sub>2</sub> O <sub>5</sub> glasses	108
7.1	Composition dependence of density for Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> and LiCl-Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> glasses	117
7.2	Composition dependence of transition temperature for Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> glasses	118
7.3	Composition dependence of molar volume for Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> and LiCl-Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> glasses	120
7.4	Composition dependence of longitudinal wave velocities for Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> and LiCl-Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> glasses	123
7.5	Composition dependence of shear wave velocities for Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> and LiCl-Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> glasses	123
7.6	Composition dependence of bulk modulus and C <sub>44</sub> for Li <sub>2</sub> O-B <sub>2</sub> O <sub>2</sub> glasses	126



77	Composition dependence of bulk modulus and C <sub>44</sub> for LiCl-Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> glasses	126
7 8	Composition dependence of Young's modulus and $C_{11}$ for $L_{12}O-B_2O_3$ glasses	127
79	Composition dependence of Young's modulus and C <sub>11</sub> for L <sub>1</sub> Cl-L <sub>12</sub> O-B <sub>2</sub> O <sub>3</sub> glasses	127
7 10	Composition dependence of Poisson's ratio for L <sub>12</sub> O-B <sub>2</sub> O <sub>3</sub> and L <sub>1</sub> Cl-L <sub>12</sub> O-B <sub>2</sub> O <sub>3</sub> glasses	129
7 11	Composition dependence of Debye temperature for Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> and LiCl-Li <sub>2</sub> O-B <sub>2</sub> O <sub>3</sub> glasses	130
8 1	Frequency dependence of ionic conductivity for (Li <sub>2</sub> O) <sub>0 1</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0 9</sub> glasses at different temperatures	134
8 2	Frequency dependence of ionic conductivity for (Li <sub>2</sub> O) <sub>0 15</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0 85</sub> glasses at different temperatures	134
8 3	Frequency dependence of ionic conductivity for $(Li_2O)_0 _2(P_2O_5)_0 _3$ glasses at different temperatures	135
8 4	Frequency dependence of ionic conductivity for (Li <sub>2</sub> O) <sub>0 25</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0 75</sub> glasses at different temperatures	135
8 5	Frequency dependence of ionic conductivity for (Li <sub>2</sub> O) <sub>0 3</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0 7</sub> glasses at different temperatures	136
8 6	Frequency dependence of ionic conductivity for (Li <sub>2</sub> O) <sub>0 35</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0 65</sub> glasses at different temperatures	136
8 7	Frequency dependence of ionic conductivity for (L <sub>12</sub> O) <sub>0 4</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0 6</sub> glasses at different temperatures	137
8 8	Frequency dependence of ionic conductivity for	137



89	Frequency dependence of ionic conductivity for $(L_{12}O)_{0.5}(P_2O_5)_{0.5}$ glasses at different temperatures	138
8 10	Frequency dependence of ionic conductivity for (L1Cl) <sub>0 1</sub> [(L1 <sub>2</sub> O) <sub>0 4</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0 6</sub> ] <sub>0 9</sub> glasses at different temperatures	138
8 1 1	Frequency dependence of ionic conductivity for (LiCl) <sub>0.2</sub> [(Li <sub>2</sub> O) <sub>0.4</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0.6</sub> ] <sub>0.8</sub> glasses at different temperatures	139
8 12	Frequency dependence of ionic conductivity for (LiCl) <sub>0.3</sub> [(Li <sub>2</sub> O) <sub>0.4</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0.6</sub> ] <sub>0.7</sub> glasses at different temperatures	139
8 13	Frequency dependence of ionic conductivity for (LiCl) <sub>0.4</sub> [(Li <sub>2</sub> O) <sub>0.4</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0.6</sub> ] <sub>0.6</sub> glasses at different temperatures	140
8.14	Frequency dependence of ionic conductivity for (LiCl) <sub>0.5</sub> [(Li <sub>2</sub> O) <sub>0.4</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0.6</sub> ] <sub>0.5</sub> glasses at different temperatures	140
8 15	Most probable site for sodium ions in sodium silicate glasses (a) Alternative sites in polar structural units. (b) Hopping within polar structural units giving rise to orientational polarization	142
8 16	Arrhenius plot of (Li <sub>2</sub> O) <sub>x</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>1-x</sub> glasses at 10 kHz	144
8 17	Arrhenius plot of (LiCl) <sub>y</sub> [(Li <sub>2</sub> O) <sub>0 4</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0 6</sub> ] <sub>1-y</sub> glasses at 10 kHz	144
8 18	Schematic illustration of asymmetric double-well potential characterizing a two-level system plotted as a function of configuration coordinate $q$ $\Delta$ is the asymmetric energy and $V$ is the average barrier height	145
8 19	Composition dependence of activation energy at 10 kHz	147
8 20	Ionic conductivity of $(L_{12}O)_x(P_2O_5)_{1-x}$ glasses at 423 K for different compositions	148
8 21	Ionic conductivity of (LiCl) <sub>y</sub> [(L <sub>12</sub> O) <sub>0.4</sub> (P <sub>2</sub> O <sub>5</sub> ) <sub>0.6</sub> ] <sub>1-y</sub> glasses at 423 K for different compositions	148



9.1	Frequency dependence of ionic conductivity for (Li <sub>2</sub> O) <sub>0.15</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.85</sub> glasses at different temperatures	153
9.2	Frequency dependence of ionic conductivity for (Li <sub>2</sub> O) <sub>0.2</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.8</sub> glasses at different temperatures	153
9.3	Frequency dependence of ionic conductivity for (Li <sub>2</sub> O) <sub>0.25</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.75</sub> glasses at different temperatures	154
9.4	Frequency dependence of ionic conductivity for (Li <sub>2</sub> O) <sub>0.3</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.7</sub> glasses at different temperatures	154
9.5	Frequency dependence of ionic conductivity for (Li <sub>2</sub> O) <sub>0.35</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.65</sub> glasses at different temperatures	155
9.6	Frequency dependence of ionic conductivity for (Li <sub>2</sub> O) <sub>0.4</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.6</sub> glasses at different temperatures	155
9.7	Frequency dependence of ionic conductivity for (LiCl) <sub>0.05</sub> [(Li <sub>2</sub> O) <sub>0.25</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.75</sub> ] <sub>0.95</sub> glasses at different temperatures	156
9.8	Frequency dependence of ionic conductivity for (LiCl) <sub>0.1</sub> [(Li <sub>2</sub> O) <sub>0.25</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.75</sub> ] <sub>0.9</sub> glasses at different temperatures	156
9.9	Frequency dependence of ionic conductivity for (LiCl) <sub>0.15</sub> [(Li <sub>2</sub> O) <sub>0.25</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.75</sub> ] <sub>0.85</sub> glasses at different temperatures	157
9.10	Frequency dependence of ionic conductivity for (LiCl) <sub>0.2</sub> [(Li <sub>2</sub> O) <sub>0.25</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.75</sub> ] <sub>0.8</sub> glasses at different temperatures	157
9.11	Frequency dependence of ionic conductivity for (LiCl) <sub>0.25</sub> [(Li <sub>2</sub> O) <sub>0.25</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.75</sub> ] <sub>0.75</sub> glasses at different temperatures	158
9.12	Arrhenius plot of (Li <sub>2</sub> O) <sub>x</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>1-x</sub> glasses at 100 Hz	159
9.13	Arrhenius plot of (LiCl) <sub>y</sub> [(Li <sub>2</sub> O) <sub>0.25</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.75</sub> ] <sub>1-y</sub> glasses at 100 Hz	160
9.14	Composition dependence of activation energy at 100 Hz	161



9.15	Ionic conductivity of (Li <sub>2</sub> O) <sub>x</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>1-x</sub> glasses at 423 K for different compositions	162
9.16	Ionic conductivity of (LiCl) <sub>y</sub> [(Li <sub>2</sub> O) <sub>0.25</sub> (B <sub>2</sub> O <sub>3</sub> ) <sub>0.75</sub> ] <sub>1-y</sub> glasses at 423 K for different compositions	163
10.1	Density and molar volume for lithium phosphate and lithium borate glasses	168
10.2	Young's modulus and bulk modulus of lithium phosphate and lithium borate glasses	168
10.3	Poisson's ratio of lithium phosphate and lithium borate glasses	170
10.4	Debye temperature of lithium phosphate and lithium borate glasses	170
10.5	Ionic conductivity of lithium phosphate glasses and lithium borate glasses	171



## LIST OF PLATES

Plate		Page
1	Alumina crucible, porcelain crucible and their lids	184
2	Electronic digital weighing machine	184
3	Furnace T1 and furnace T2	185
4	Metal tongs	185
5	Dessicator	186
6	Low speed diamond cutter	186
7	Digital vernier caliper	187
8	Sample holder for ultrasonic measurement	187
9	Configuration of LCR meter, electric furnace and personal computer for ionic conductivity measurement	188
10	HP 16809B Kelvin Clip Leads and sample holder for ionic conductivity measurement	188



### LIST OF ABBREVIATIONS

Mol% Mole percent

fcc Face-centered cube

T-T-T Time-Temperature-Transformation Curve

NMR Nuclear Magnetic Resonance

NBO Non-Bridging Oxygen

BO Bridging Oxygen

NDT Non-Destructive Testing

RF Radio Frequency

DUT Device Under Test

IR Infrared

ADWP Asymmetric Double Well Potential



Abstract of the thesis presented to the Senate of Universiti Putra Malaysia in fulfilment of the requirements for the degree of Master of Science.

## THE STUDIES OF ELASTIC PROPERTIES AND IONIC CONDUCTIVITY OF LITHIUM PHOSPHATE AND LITHIUM BORATE GLASSES

By

### **LOW YEE SAN**

June 1999

Chairman : Associate Professor Sidek Hj. Abd. Aziz (Ph.D.)

Faculty : Science and Environmental Studies

Glasses are of scientific interest because their compositions can be varied over a very wide range and it is possible to produce glasses for a particular application. In this work, lithium glasses with phosphate and borate as the glass formers are studied. Four series of lithium glasses, lithium phosphate  $(Li_2O)_x(P_2O_5)_{1-x}$ (x 0.10 to 0.50), lithium chlorophosphate  $(\text{LiCl})_{y}[(\text{Li}_{2}O)_{0.4}(P_{2}O_{5})_{0.6}]_{1-y}$  (y = 0.10 to 0.50), lithium borate  $(\text{Li}_{2}O)_{x}(B_{2}O_{3})_{1-x}$  (x = 0.15 to 0.40) and lithium chloroborate  $(LiCl)_{v}[(Li_{2}O)_{0.25}(B_{2}O_{3})_{0.75}]_{1-y}$  (y = 0.05 to 0.25) glasses have been prepared for the elastic property and ionic conductivity measurement using the rapid quenching technique. The ultrasonic velocities of the glass samples were measured by using the MBS 8000 ultrasonic data acquisition system. The densities were determined using the Archimedes' principle. The elastic properties and Debye temperatures of these glasses were then obtained. Young's modulus, bulk modulus and Debye temperature of binary lithium phosphate and borate glasses were found to increase with the mole fraction of Li<sub>2</sub>O. The elastic moduli and Debye temperature of lithium chlorophosphate glasses were found decrease with the addition of LiCl whereas anomalous trend of elastic moduli and Debye temperature were observed in the case of lithium chloroborate glasses. The elastic properties of these glasses are closely related to the strength of glass networks and structures. The ionic conductivities of these glasses were measured by using the HP LCR meter. The conductance G was measured and used to calculate the ionic conductivity. Unusual frequency-dependent ionic conductivity was encountered in these glasses at lower mole fractions of Li<sub>2</sub>O or LiCl content. Anomalous trend of temperature-dependent ionic conductivity was observed for binary glasses with lower mole fractions of alkali content. Generally, the ionic conductivity of these glasses was found to increase with the Li<sub>2</sub>O or LiCl content.



Abstrak tesis yang dikemukakan kepada Senat Universiti Putra Malaysia sebagai memenuhi keperluan untuk ijazah Master Sains.

### KAJIAN CIRI-CIRI KENYAL DAN KEKONDUKSIAN IONIK UNTUK KACA JENIS LITIUM FOSFAT DAN LITIUM BORAT

### Oleh

### **LOW YEE SAN**

Jun 1999

Pengerusi : Profesor Madya Sidek Hj. Abd. Aziz (Ph.D.)

Fakulti : Sains dan Pengajian Alam Sekitar

Kaca mempunyai kepentingan saintifik kerana ia dapat dihasilkan dalam pelbagai komposisi untuk aplikasi tertentu. Dalam kajian ini, kaca litium dengan fosfat dan borat sebagai pembentuk kaca telah dikaji. Empat siri kaca litium telah dihasilkan iaitu lithium fosfat (Li<sub>2</sub>O)<sub>x</sub>(P<sub>2</sub>O<sub>5</sub>)<sub>1-x</sub> (x = 0.10 hingga 0.50), litium klorofosfat (LiCl)<sub>y</sub>[(Li<sub>2</sub>O)<sub>0.4</sub>(P<sub>2</sub>O<sub>5</sub>)<sub>0.6</sub>]<sub>1-y</sub> (y = 0.10 hingga 0.50), litium borat (Li<sub>2</sub>O)<sub>x</sub>(B<sub>2</sub>O<sub>3</sub>)<sub>1-x</sub> (x = 0.15 hingga 0.40) dan litium kloro-borat (LiCl)<sub>y</sub>[(Li<sub>2</sub>O)<sub>0.25</sub>(B<sub>2</sub>O<sub>3</sub>)<sub>0.75</sub>]<sub>1-y</sub> (y = 0.05 hingga 0.25) menerusi teknik penyejukan mendadak. Pengukuran halaju ultrasonik telah dijalankan dengan menggunakan sistem pemperolehan data ultrasonik MBS 8000. Ketumpatan kaca-kaca diukur dengan mengunakan prinsip Archimedes. Daripada halaju ultrasonik dan



ketumpatan, ciri-ciri kenyal kaca-kaca ini dihitung. Modulus Young, modulus pukal dan suhu Debye untuk kaca litium fosfat dan litium borat didapati meningkat dengan kandungan Li<sub>2</sub>O atau LiCl. Modulus kenyal dan suhu Debye untuk kaca litium kloro-fosfat berkurangan dengan penambahan LiCl manakala corak anomalus bagi modulus kenyal dan suhu Debye telah diperhatikan dalam kes kaca litium kloro-borat. Ciri-ciri kenyal kaca-kaca ini didapati berkait rapat dengan kekuatan struktur kaca. Kekonduksian ionik bagi kaca-kaca ini telah diukur dengan menggunakan mesin meter pengukuran HP LCR. Konduktan G telah diukur untuk semua sampel kaca dan diguna untuk pengiraan kekonduksian ionik. Pemerhatian yang agak luar biasa didapati untuk hubungan antara frekuensi dan kekonduksian ionik pada kandungan Li<sub>2</sub>O atau LiCl yang rendah. Corak anomalus telah diperhati untuk kekonduksian ionik yang bersandarkan suhu untuk kaca litium fosfat dan litium borat pada kandungan alkali yang rendah. Umumnya, kekonduksian ionik bagi kaca-kaca ini meningkat dengan penambahan Li<sub>2</sub>O atau LiCl.



#### **CHAPTER I**

### RESEARCH OVERVIEW

### Introduction

Glasses are among the most ancient materials in human civilisation but the strong interaction between fundamental research and glass technology appeared only in the middle of this century. The reason is perhaps that the glassy materials are out of thermodynamic equilibrium and that their physical properties depend on their thermal histories. In addition, classical theories about crystal structures and classical investigation technique seem to be ineffective in the studies of these kinds of disordered structures.

In spite of the difficulties, glasses are still of scientific interest because their compositions can be varied over a very wide range and it is possible to produce glasses for particular applications. As our knowledge about glasses is growing, glasses have found a lot of applications of technological importance such as optical wave guides, glasses for lasers, amorphous semiconductors in xerography and solar cells or electrolytes in high density batteries.



In this research, lithium glasses with phosphate and borate as the glass formers are studied. Lithium glasses are well known for their ionic conductivity but the elastic properties of this type of glasses have not gained much interest from researchers. A brief review of the ionic conductivity and elastic properties of these glasses will first be described in this chapter. The objective of this work and the chapter organisation will be discussed subsequently.

### Literature Reviews

There are numerous publications on the ionic conductivity of lithium glasses but few are found on the ultrasonic properties. Most of the researchers are more interested on borate glasses rather than phosphate glasses.

### Phosphate Glasses

The hygroscopic behaviour of crystalline and vitreous phosphate is one of the reasons for little investigations on the structure and the properties of these glasses. However, phosphate-based glasses still contribute to some important technological applications, such as bioceramic, glass to metal seals, fast ionic conducting electrolytes, high quality micro-optic lenses and host material for rare-earth solid state laser. In addition, a lead-iron phosphate glass is used as a stable medium for storage of high level nuclear waste.

In the study of lithium phosphate glasses, Martin and Angell (1986) have studied the electrical conductivity of lithium phosphate glasses over a wide

