

STRUCTURAL, ELASTIC AND OPTICAL PROPERTIES OF ZINC-ALUMINO-BOROSILICATE DOPED GADOLINIUM OXIDE GLASS-CERAMICS



NUR ATIKAH NAZIHAH BINTI ISMAIL

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DEDICATIONS

To my loving parents, For their unconditional love and support To accomplish my research with confidence

> To my siblings and family For making my life complete

To my husband and his family For constant source of support

To all my very wonderful friends For making my life full of joy and happiness

To all my lecturers For helping and guiding me to complete this study

Thank you all

Abstract of thesis presented to the Senate of Universiti Putra Malaysia in fulfilment of the requirement for the degree of Master of Science

STRUCTURAL, ELASTIC AND OPTICAL PROPERTIES OF ZINC-ALUMINO-BOROSILICATE DOPED GADOLINIUM OXIDE GLASS-CERAMICS

By

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In recent years, researchers are developing a great interest towards the fabrication and synthesizing willemite (Zn_2SiO_4) glass-ceramics. However, intensity luminescence of willemite based glass ceramic doped with gadolinium oxide (Gd_2O_3) is less reported. Hence, this study is focusing on fabricate and synthesized willemite based glass-ceramics using Gd_2O_3 as dopant. The Gd_2O_3 doped zinc-alumino-borosilicate $(ZnO-Al_2O_3-B_2O_3-SiO_2)$ glass system were synthesized via the melt-quenching approach using composition $(60-x)ZnO-5Al_2O_3-15B_2O_3-20SiO_2:x(Gd_2O_3)(x = 0, 0.5, 1.0, 2.0 and 3.0 mol%) and willemite-based glass-ceramics were obtained from these precursor glasses through a controlled crystallization process.$

According to all analysis, the best willemite based glass-ceramic has been selected from the glass system when doped with 3 mol% of Gd₂O₃ and temperature at 800 °C. XRD and FESEM methods were also used to analyze the structural properties of precursor glass, the formation of willemite crystal phase, shape, and size as crystallization temperatures increased. The average approximated crystallite size determined by XRD is in the 50-70 nm range. The structural properties of glass and glass-ceramics were assessed using FTIR spectroscopy. The infrared spectra studies reveal the existence of SiO₂ and ZnO₄ vibrational groups indicating the establishment of the willemite crystal phase. As for elastic analysis, the values for experimental elastic moduli were obtained from ultrasonic velocities measurement by using the non-destructive ultrasonic technique. The longitudinal and shear velocities vary from 4798 to 6976 m/s and 2991 to 3082 m/s, respectively. The experimental elastic moduli (longitudinal modulus (L), shear modulus (G), bulk modulus (K) and Young's modulus (E)) increases from 94.60 to 202.98 GPa, 36.77 to 39.61 GPa, 45.56 to 150.17 GPa and 86.94 to 109.22 GPa, respectively. In addition, the optical band willemite based glass-ceramics doped with Gd₂O₃ fluctuate from 3.64 to 3.38 eV due to structural rearrangement of network. The emission spectra of gadolinium ions show the strong emission peak at wavelengths of 425, 447, 462, 485, and 530

nm. The willemite phase shows prominent green emission spectra located at 530 nm. These emission spectra produce when the dopant content and heat treatment temperatures increase, the luminescence performance of the glass-ceramics also improves. The incorporation of gadolinium ions into the willemite crystals as heat treatment temperatures increase affect the intensity of the emission. The structural, elastic and optical properties of willemite glass-ceramics enhanced with the addition of Gd_2O_3 as dopant.



Abstrak tesis yang dikemukakan kepada Senat Universiti Putra Malaysia sebagai memenuhi keperluan untuk ijazah Master Sains

SIFAT STRUKTUR, ELASTIK DAN OPTIK ZINK-ALUMINA-BOROSILIKA DOP GADOLINIUM OKSIDA KACA-SERAMIK

Oleh

NUR ATIKAH NAZIHAH BINTI ISMAIL

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Beberapa tahun kebelakangan ini, penghasilan dan kajian terhadap willemite (Zn₂SiO₄) berasaskan kaca-seramik telah berjaya menarik minat dan tumpuan ramai penyelidik. Walau bagaimanapun, keamatan pendarkilau wilemit berasaskan kaca-seramik didopkan dengan gadolinium oksida Gd_2O_3 kurang dilaporkan. Oleh itu, kajian ini menumpukan untuk fabrikasi dan penghasilan wilemit berasaskan kaca-seramik dengan menggunakan Gd_2O_3 sebagai bahan pengedop. Sistem kaca zink-alumina-borosilika (ZnO-Al₂O₃-B₂O₃-SiO₂) didopkan dengan Gd_2O_3 telah dihasilkan melalui kaedah sepuh lindap menggunakan komposisi (60-x)ZnO- $5Al_2O_3-15B_2O_3-20SiO_2$: $x(Gd_2O_3)(x = 0, 0.5, 1.0, 2.0 dan 3.0 mol%)$ dan wilemit berasaskan kaca-seramik diperolehi dari bahan pemula kaca melalui proses penghabluran yang dikawal.

Menurut semua analisis, wilemit berasaskan kaca-seramik terbaik telah dipilih daripada sistem kaca apabila didopkan dengan 3 mol% Gd₂O₃ dan bersuhu pada 800 °C. Kaedah XRD dan FESEM digunakan untuk menganalisis sifat struktur bahan pemula kaca, penghasilan fasa hablur wilemit, bentuk dan saiz apabila suhu proses penghabluran meningkat. Purata anggaran saiz hablur halus ditentukan menggunakan XRD adalah dalam julat 50-70 nm. Sifat struktur kaca dan kaca-seramik dinilai menggunakan FTIR spektroskopi. Kajian spektrum inframerah mendedahkan kewujudan kumpulan getaran SiO2 dan ZnO4 menunjukkan pertumbuhan fasa hablur wilemit. Manakala untuk analisis elastik, nilai uji kajian modulus kenyal diperoleh daripada pengukuran halaju gelombang yang menggunakan teknik tak musnah ultrasonik. Halaju membujur dan ricih berubah dari 4798 ke 6976 m/s dan 2991 ke 3082 m/s masing-masing. Uji kajian modulus kenyal (modulus membujur (L), modulus ricih (G), modulus pukal (K) and modulus Young (E)) meningkat dari 94.60 ke 202.98 GPa, 36.77 ke 39.61 GPa, 45.56 ke 150.17 GPa dan 86.94 ke 109.22 GPa masing-masing. Tambahan pula, tenaga jurang jalur kaca dan kaca-seramik berkurang dari 3.64 ke 3.38 eV oleh kerana penyusunan semula struktur rangkain. Spektrum pancaran gadolinium ion menunjukkan puncak pancaran kuat di panjang gelombang 425, 447, 462, 485 dan 530 nm. Fasa wilemit menunjukkan pancaran utama berwarna hijau dilokasi 530 nm. Spektrum pancaran tersebut dihasilkan apabila kandungan bahan pengedop dan suhu rawatan haba meningkat, prestasi pendarkilau kaca-seramik juga bertambah baik. Penerapan gadolinium ion ke dalam hablur wilemit apabila suhu rawatan haba meningkat mempengaruhi keamatan pancaran. Sifat struktur, elastik dan optik willemite berasaskan kaca-seramik dipertingkatkan dengan penambahan Gd₂O₃ sebagai bahan pengedop.



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LIST OF ABBREVIATIONS

Symbols	Description
Al ₂ O ₃	Aluminium oxide
B ₂ O ₃	Boron trioxide
ВаО	Barium oxide
CaO	Calcium oxide
DSC	Differential scanning calorimetry
Eg	Optical band gap
E	Young's modulus
FTIR	Fourier transform infrared
FESEM	Field emission scanning electron microscopy
Gd ³⁺	Gadolinium ions
Gd ₂ O ₃	Gadolinium oxide
G	Shear modulus
JCPDS	Joint committee on powder diffraction standards
K ₂ O	Potassium oxide
к	Bulk modulus
L	Longitudinal modulus
MgO	Magnesium oxide
Na ₂ O	Sodium oxide
PLD	Pulsed laser deposition
PL	Photoluminescence
SiO ₂	Silica oxide
UV	Ultraviolet
UV-Vis	Ultraviolet-Visible

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Vm	Molar volume
V _L	Longitudinal ultrasonic velocity
V _S	Shear ultrasonic velocity
XRD	X-Ray diffraction
Zn ₂ SiO ₄	Willemite
Zn ₂ SiO ₄ :Gd ³⁺	Willemite doped Manganese Oxide
ZABS	Zinc-alumino-borosilicate
ZnO	Zinc oxide
α	Alpha
β	Beta
Y	Gamma
ρ	Density
σ	Poisson's ratio

CHAPTER 1

INTRODUCTION

1.1 Research background

In recent years, researchers are developing a great interest toward transparent glass and glass-ceramics due to availability as attractive materials in optical devices. Transparent glass-ceramics have great shaping versatility of glasses and the optical efficiency of crystals (Yaowakulpattana et al., 2015). Furthermore, due to the outstanding optical properties for neon discharge lamps, fluorescent lamps, color TVs, light emitting diodes, oscilloscopes, and other displays and lighting devices, the manufacturing and analysis of glass and glass-ceramics based phosphors has been explored (Bernardo et al., 2008; Takesue et al., 2009; Sahu et al., 2016).

Transition metal and rare earth ions doped glass and glass ceramics are interesting optical host materials because they boost the optical properties of the system by exhibiting high transparency from the ultraviolet and infrared regions, a low refractive index, thermal stability, and a high likelihood of assimilating a large number of dopant ions (Hou et al., 2014). Several studies on synthesize of willemite glass-ceramics doped with europium ions and samarium ions were reported using sintering process (Tarafder et al., 2013; Tarafder et al., 2014). From the research study, it was observed the rare earth added improving the optical quality of the material as the luminescence of europium ions and samarium ions showing emission transitions from ${}^{5}D_{0} \rightarrow {}^{7}F_{4}$ and ${}^{4}G_{5/2} \rightarrow {}^{6}H_{11/2}$ energy state. According to the result, it was discovered red emission spectra based on absorption band around 601 and 615 nm. Recently, Effendy and coworkers synthesize and studies the optical performance of europium ions doped willemite glass-ceramic (Effendy et al., 2019). From the results, they observed strong green emission peak around 557 nm with three different transition states of europium ions photon excitation. To the best of our knowledge, no reports on the luminescence performance of gadolinium ions doped willemite glass-ceramic have been published. We discovered that among numerous rare earth ions, gadolinium ions have emerged as a superior candidate for optical host materials because it emits intense green fluorescence in the visible range. Furthermore, previous work also states that the gadolinium ions doped materials attract greater interest in lighting industry for various applications (Abo-Naf et al., 2015).

Zinc silicate (Zn₂SiO₄), known colloquially as willemite, a zinc ore mineral with a phenakite structure, is one of the most common zinc ore minerals. Willemite persists in multiple polymorphs, crystalizing in various space groups (Akimoto, 1967). Willemite powder is produced using a variety of standard processes, including solid-state diffusion, pulsed laser deposition (PLD), sol-gel forced precipitation, organometallic complex route, dry reaction, spray-pyrolysis,

combustion methods, polymer aided and hydrothermal procedures (Su et al., 1996; Bhatkar, 2007; Sharma & Bhatti, 2011; Van & Klement, 2016). However, the most preferred technique to obtain willemite is melt-quenching method because it is simple, homogeneous and easy shapes formation (Abdel-hameed & Margha, 2020). Additionally, the excellent properties of willemite, such as chemical stability, transparency in the visible range, and ultraviolet transparency, will contribute to enhance the optical characteristics that meet the criteria for phosphor materials in a variety of applications, such as televisions, fluorescent lamps, and other lighting or display devices (Brunold et al., 1996; Chakradhar et al., 2004; Wan et al., 2006; Yan & Huang, 2007; Takesue et al., 2009;).

In the proposed research, a series of gadolinium ions doped zinc-aluminoborosilicate (ZABS) glasses are made using a standard melt-quenching approach. The gadolinium ions doped willemite glass-ceramics are obtained by a controlled heat treatment procedure from these precursor glasses. Characterization of the precursor glass and willemite based glass-ceramics in terms of structural, elastic, and optical performance were conducted to evaluate the influence of heat treatment temperature and Gd₂O₃ doping. X-ray diffraction (XRD) and Fourier transform infrared (FTIR) were employed to study the structural characteristics of precursor glass and gadolinium ions doped willemite glass ceramics at varying heat treatment temperatures, while ultrasonic velocity was used to study the elastic properties. The optical properties were studied using UV-Visible (UV-Vis) and photoluminescence (PL) spectroscopy. Thus, the focus of this research is to fabricate and characterize gadolinium ions doped willemite glass-ceramics in a zinc-alumino-borosilicate (ZnO-Al₂O₃-B₂O₃-SiO₂) glass system.

1.2 Problem statement

Previously, willemite glass-ceramics materials containing transition metal and rare earth ions have been encounter greater emphasis. Rare-earth doped willemite glass-ceramics are piquing the scrutiny of scientists in laser cooling, solid-state lasers, optical devices for three-dimensional color displays, optical communications, and upconverting optical components (Tick et al., 1995; Sohn et al., 1999; Suzuki & Ohishi, 2004; Campbell et al., 2011; Nemova & Kashyap, 2012). Given the fact that willemite glass-ceramics is significant with high chemical and physical stability and high luminescence efficiency (El Ghoul et al, 2012c). Generally, willemite glass-ceramics as luminescent materials produce emission of light due to absorption of energy. Impurities and defects in crystals obtained in willemite glass-ceramics causing luminescence instability due to low energy absorption rates. In order to obtain high intensity luminescence, willemite glass-ceramics have high possibility of being a multi-color phosphor by doped with various rare earth ions producing emission ranging from ultraviolet to infrared spectral range (El Ghoul, 2018). Several reports on luminescence performance of willemite glass-ceramics doped with different rare earth ions such as manganese ions for green emission, gallium ions for violet emission, cerium ions for blue emission and europium ions for red emission color (Lu & Yun, 2013; El Ghoul, 2018; Nuraidayani et al., 2019). However, to the best of our

knowledge, no reports on the luminescence performance of gadolinium ions doped willemite glass-ceramic have been published. We discovered that among numerous rare earth ions, gadolinium ions have emerged as a superior candidate for optical host materials because it emits intense green fluorescence in the visible range. Besides, Gd³⁺ increase energy transfer efficiency towards Ce³⁺, Tb³⁺ and Mn³⁺ luminescent center that act as a sensitizer in scintillator or phosphor materials (Martino et al., 2008). Furthermore, previous work also states that the gadolinium ions doped materials attract greater interest in lighting industry for various applications (Abo-Naf et al., 2015).

Usually, band-gap energy of willemite glass is related with electronic band structure that depending on the temperature (Sessolo & Bolink, 2011). Energy bandgap tend to decreased with progression of temperatures leading to increase in atomic vibrations and resulting to widen average interatomic spacing (Alibe et al., 2021). Generally, low green emission color and high optical band gap of zinc silicate glass were overcome by heat treatment process to produce multiple emissions with wide optical band gap in willemite glass-ceramics. To address this issue, ZnO-Al₂O₃-B₂O₃-SiO₂ (ZABS) glass systems doped with Gd₂O₃ are synthesized followed by controlled heat treatment process to study their influence on the properties of willemite glass ceramics.

In the proposed research, a series of gadonium doped zinc aluminium borosilicate (ZABS) glasses are made using a standard melt-quenching approach. The Gd³⁺-doped willemite glass-ceramics are obtained by a controlled heat treatment procedure from these precursor glasses. Characterization of the precursor glass and willemite based glass-ceramics in terms of structural, elastic, and optical performance were conducted to evaluate the influence of heat treatment temperature and Gd₂O₃ doping. X-ray diffraction (XRD) and Fourier transform infrared (FTIR) were employed to study the structural characteristics of precursor glass and Gd³⁺-doped willemite glass ceramics at varying heat treatment temperatures, while ultrasonic velocity was used to study the elastic properties. The optical properties were studied using UV-Visible (UV-Vis) and photoluminescence (PL) spectroscopy. Thus, the focus of this research is to fabricate and characterize Gd³⁺-doped willemite glass-ceramics in a ZnO-Al₂O₃-B₂O₃-SiO₂ glass system.

1.3 Objectives of the study

The primary goal of this research is to build and synthesize willemite-based glass-ceramics doped with gadolinium oxide. This project entails the identification of glass with various compositions, the melt-quenching method, a variety of heat treatment processes, the advancement of the doping process and fundamental studies of the crystallization process.

1. To synthesize the Gd³⁺-doped willemite glass-ceramics derived from ZnO-Al₂O₃-B₂O₃-SiO₂ glass system.

- To study the effect Gd₂O₃ doping on structural, elastic and optical properties of precursor glass and willemite glassceramics.
- To analyze the influence of various heat treatment temperatures on structural, elastic and optical properties of precursor glass, willemite based glass-ceramics and Gd³⁺-doped willemite glassceramics.

1.4 Scopes of the study

The study's scopes are as follows in order to achieve the study's goal:

- A series of precursor glass based on stoichiometric equation of (60-x)ZnO-5Al₂O₃-15B₂O₃-20SiO₂-x(Gd₂O₃) where x = 0, 0.5, 1.0, 2.0 and 3.0 mol% has been prepared using ZnO, Al₂O₃, B₂O₃, SiO₂ and Gd₂O₃ powder by simple melt-quenching technique.
- 2) The glass transition temperature (T_g) and glass crystallization temperature (T_c) have been measured using DSC spectroscopy.
- 3) Willemite based glass-ceramics has been derived from the precursor ZnO-Al₂O₃-B₂O₃-SiO₂ (ZABS) glass system by a controlled crystallization process.
- 4) The structural, elastic and optical properties of precursor ZABS glass, willemite glass-ceramics and Gd³⁺-doped willemite glass ceramics has been analyzed using Archimedes method, XRD, FTIR, FESEM, ultrasonic velocity, UV-Vis and PL spectroscopy.

1.5 Hypothesis

Hypothesis of the study are related with gadolinium ions that improving the properties of network. Gadolinium ions has many strong absorption bands that emit one of strong green fluorescence and thus, improving the optical properties of glass and glass-ceramics. Therefore, incorporation of gadolinium ions in ZABS glass systems will improve the optical properties of glass and willemite glass-ceramics.

1.6 Importance of the study

Nowadays, due to excellent luminous materials, many glass systems have a wide range of applications in a variety of sectors. To optimize the structural, elastic, and optical characteristics of transparent glass systems and glass-ceramics, transition metal (TM) and rare earth (RE) ions are doped into the systems. A vast amount of research is focusing on the production and synthesizing of glass and glass-ceramics based phosphor doped with gadolinium ions due to possible applications in lighting industry.

Willemite is a prominent and frequently used phosphor in optic, optoelectronic, and lighting technologies. All of the atoms are in the same general location and are made up of a tetrahedral framework. As a result, willemite is recognized as a good host matrix for several TM and RE ions in order to achieve high efficiency luminescence (Tarafder et al., 2014).

There have been few recent publications on the structural, elastic, and optical characteristics of Gd³⁺-doped willemite glass-ceramics generated from precursor glass. The current study reports on the production of a precursor ZnO-Al₂O₃-B₂O₃-SiO₂ glass system using a traditional melt-quenching approach and the derived willemite-based glass-ceramics using a control crystallization process of precursor glasses. Thus, the willemite glass-ceramics has been doped with gadolinium oxide (Gd₂O₃) to increase the properties and quality of the final products.

1.7 Outline of thesis

The thesis consists of five chapters and the arrangement is structured as follows. Chapter 1 gives an introduction of research background that consists of willemite glass ceramics and Gd^{3+} -doped willemite glass-ceramics, the problem statements, the objectives, the scopes and also the importance of this study. Chapter 2 discussed glass theory, glass-ceramics, and earlier studies, both past and present, that have been carried out by other researchers as a guideline and citations. The list of substances, apparatus, method to prepare the samples and characterization of the precursor glass, willemite glass-ceramics and willemite glass-ceramics doped gadolinium oxide are elaborated in Chapter 3. Chapter 4 is the most significant part as the result concerning the effect of different percentage Gd_2O_3 doping content and progression of heat treatment temperatures towards the structural, elastic and optical properties of precursor glass $ZnO-Al_2O_3-B_2O_3-SiO_2$ and Gd^{3+} -doped willemite glass-ceramics are analyzed and discussed in Chapter 5.

REFERENCES

- Abdel-Hameed, S. A. M., Abo-Naf, S. M., & Hamdy, Y. M. (2019a). The effect of heat treatment on photoluminescence and magnetic properties of new yellow phosphor based on sanbornite (BaSi₂O₅) glass ceramic doped with Gd³⁺ and Mn²⁺. *Journal of Non-Crystalline Solids*, *517*, 106–113.
- Abdel-Hameed, S. A. M., Hamdy, Y. M., & Sadek, H. E. H. (2019b). Characterization and luminescence properties of Mn-doped zinc borosilicate glasses and glass-ceramics. *Silicon*, *11*(3), 1185–1192.
- Abdel-hameed, S. A. M., & Margha, F. H. (2020). Preparation, crystallization and photoluminescence properties of un-doped nano willemite glass ceramics with high ZnO additions. *Optik - International Journal for Light and Electron Optics 206*, 164374.
- Abo-Naf, S. M., Abdel-Hameed, S. A. M., Marzouk, M. A., & Elwan, R. L. (2015). Sol–gel synthesis, paramagnetism, photoluminescence and optical properties of Gd-doped and Bi–Gd-codoped hybrid organo-silica glasses. *Journal of Materials Science: Materials in Electronics*, 26(4), 2363–2373.
- Abo-Naf, S. M., & Marzouk, M. A. (2021). Tunable blue and green emissions of sol–gel synthesized transparent nano-willemite codoped with nano-pyroxmangite and dysprosium. *Nano-Structures and Nano-Objects*, *26*, 100685.
- Akimoto, Y. A. S. and S. (1967). High pressure transformations in zinc germanates and silicates. *Journal of Solid State Chemistry*, *215*(5108), 1367–1368.
- Ali, E. A. G. E., Matori, K. A., Saion, E., Aziz, S. H. A., Zaid, M. H. M., & Alibe, I. M. (2018). Calcination effect to the physical and optical properties of Zn₂SiO₄ composite prepared by impregnation of ZnO on SiO₂ amorphous nanoparticles. *IOP Conference Series: Materials Science and Engineering*, 440(1), 75–85.
- Alibe, I. M., Matori, K. A., Sidek, H. A. A., Yaakob, Y., Rashid, U., Alibe, A. M., Zaid, M. H. M., & Khiri, M. Z. A. (2018). Effects of calcination holding time on properties of wide band gap willemite semiconductor nanoparticles by the polymer thermal treatment method. *Molecules*, 23(4), 1–18.
- Alibe, I. M., Matori, K. A., Zaid, M. H. M., Nasir, S., Alibe, A. M., & Khiri, M. Z. A. (2021). Polymer thermal treatment production of cerium doped willemite nanoparticles: An analysis of structure, energy band gap and luminescence properties. *Materials*, 14(5), 1–26.
- Altaf, M., Chaudhry, M. A., & Zahid, M., (2003). Study of optical band gap of zincborate glasses. *Journal of Research (Science)*, *14*(2), 253–259.

- Anigrahawati, P., Sahar, M. R., & Ghoshal, S. K. (2015). Influence of Fe₃O₄ nanoparticles on structural, optical and magnetic properties of erbium doped zinc phosphate glass. *Materials Chemistry and Physics*, *155*, 155–161.
- Anupama, M. K., & Rudraswamy, B. (2016). Effect of Gd³⁺- Cr³⁺ ion substitution on the structural, electrical and magnetic properties of Ni - Zn ferrite nanoparticles. *IOP Conference Series: Materials Science and Engineering*, *149*(1), 012194.
- Aronne, A., Esposito, S., & Pernice, P. (1999). FTIR and DTA study of structural transformations and crystallisation in BaO - B₂O₃ - TiO₂ glasses. Society of *Glass Technology*, 40(2), 63-68(6).
- Babu, B. C., Naresh, V., Prakash, B. J., & Buddhudu, S. (2011). Structural, thermal and dielectric properties of lithium zinc silicate ceramic powders by sol-gel method. *Ferroelectrics, Letters Section*, 38(4–6), 114–127.
- Baer, D. R., & Thevuthasan, S. (2010). Chapter 16 Characterization of thin films and coatings. In *Handbook of Deposition Technologies for Films and Coatings* (Third Edition) (pp. 749-864). William Andrew Applied Science Publishers.
- Baghramyan, V. V., Sargsyan, A. A., Knyzyan, N. B., Harutyunyan, V. V., Badalyan, A. H., Grigoryan, N. E., Aprahamian, A. A., & Manukyan, K. V. (2020). Pure and cerium-doped zinc orthosilicate as a pigment for thermoregulating coatings. *Ceramics International*, 46(4), 4992–4997.
- Bahari, H., Hj, S., Aziz, A., & Kamari, H. M. (2012). The effect of bismuth on the structure and mechanical properties of GeO₂ – PbO – Bi₂O₃ ternary bulk glass system. *Journal of the Ceramic Society of Japan, 120,* 280-285.
- Beall, G. H., & Duke, D. A. (1969). Transparent glass-ceramics. Journal of Materials Science, 4(4), 340–352.
- Bernardo, E., Esposito, L., Rambaldi, E., Tucci, A., & Hreglich, S. (2008). Recycle of waste glass into "glass-ceramic stoneware." *Journal of the American Ceramic Society*, 91(7), 2156–2162.
- Bhatkar, V. B. (2007). Combustion synthesis of silicate phosphors. Optical Materials, 29, 1066–1070.
- Boivin, D., Thomas, F., Burov, E., Pastouret, A., Gonnet, C., Cavani, O., Collet, C., & Lempereur, S. (2010). Quenching investigation on new erbium doped fibers using MCVD nanoparticle doping process. *7580*(0), 1–9.
- Bootjomchai, C. (2015). Comparative studies between theoretical and experimental of elastic properties and irradiation effects of soda lime glasses doped with neodymium oxide. *Radiation Physics and Chemistry*, *110*, 96–104.

- Bridge, B., Patel, N. D., & Waters, D. N. (1983). On the elastic constants and structure of the pure inorganic oxide glasses. *Physica Status Solidi (a), 77,* 655-668.
- Brunold, T. C., Güdel, H. U., & Cavalli, E. (1996). Absorption and luminescence spectroscopy of Zn₂SiO₄ willemite crystals doped with Co²⁺. *Chemical Physics Letters*, 252(1–2), 112–120.
- Campbell, J. H., Hayden, J. S., & Marker, A. (2011). High-power solid-state lasers: A laser glass perspective. *International Journal of Applied Glass Science*, *2*(1), 3–29.
- Casasola, R., Rincón, J. M., & Romero, M. (2012). Glass-ceramic glazes for ceramic tiles: A review. *Journal of Materials Science*, 47(2), 553–582.
- Cetinkaya Colak, S., Akyuz, I., & Atay, F. (2016). On the dual role of ZnO in zincborate glasses. *Journal of Non-Crystalline Solids*, *432*, 406–412.
- Chakradhar, R. P. S., Nagabhushana, B. M., Chandrappa, G. T., Ramesh, K. P., & Rao, J. L. (2004). Solution combustion derived nanocrystalline Zn₂SiO₄:Mn phosphors: A spectroscopic view. *Journal of Chemical Physics*, *121*(20), 10250–10259.
- Chandra Babu, B., Rao, B. V., Ravi, M., & Babu, S. (2017). Structural, microstructural, optical, and dielectric properties of Mn²⁺: Willemite Zn₂SiO₄ nanocomposites obtained by a sol-gel method. *Journal of Molecular Structure*, *1127*, 6–14.
- Chatterjee, A. K. (2001). Handbook of analytical techniques in concrete science and technology. *Handbook of Analytical Techniques in Concrete Science and Technology*, 275–332.
- Chen, H., Ding, J., Guo, W., Chen, G., & Ma, S. (2013). Blue-green emission mechanism and spectral shift of Al-doped ZnO films related to defect levels. *RSC Advances*, *3*(30), 12327–12333.
- Chua, E., & Kamari, H. M. (2016). Elastic properties of TeO₂-B₂O₃-ZnO-Gd₂O₃ glasses using non-destructive ultrasonic technique. *Chalcogenide Letters*, *13*, 281-289.
- Coppola, V., Boni, M., Gilg, H. A., Balassone, G., & Dejonghe, L. (2008). The "calamine" nonsulfide Zn-Pb deposits of Belgium: Petrographical, mineralogical and geochemical characterization. *Ore Geology Reviews*, 33(2), 187–210.
- Deboer, B. G., & Berkowitz, J. K. (1991). First principles investigation of electronic structure and associated properties of zinc orthosilicate phosphors. *Journal of Luminescence*, *47*, 197–206.

- Devi, A. V. G., Rajendran, V., & Rajendran, N. (2010). Ultrasonic characterisation of calcium phosphate glasses and glass-ceramics with addition of TiO₂. *International Journal of Engineering Science and Technology*, 2(6), 2483–2490.
- Dhananjaya, N., Nagabhushana, H., Nagabhushana, B. M., Rudraswamy, B., & Sharma, S. C. (2012). Effect of different fuels on structural, thermo and photoluminescent properties of Gd₂O₃ nanoparticles. *Spectrochimica Acta Part A: Molecular And Biomolecular Spectroscopy*, *96*, 532–540.
- Dong, G., Tian, G., Gong, L., Tang, Q., Li, M., Meng, J., & Liang, J. (2020). Mesoporous zinc silicate composites derived from iron ore tailings for highly efficient dye removal: Structure and morphology evolution. *Microporous* and Mesoporous Materials, 305, 110352.
- Effendy, N., Ab Aziz, S. H., Mohamed Kamari, H., Mohd Zaid, M. H., & Abdul Wahab, S. A. (2020). Ultrasonic and artificial intelligence approach: Elastic behavior on the influences of ZnO in tellurite glass systems. *Journal of Alloys and Compounds*, *835*, 1–11.
- Effendy, N., Aziz, S. H. A., Kamari, H. M., Matori, K. A., & Zaid, M. H. M. (2019). Enhanced green photoluminescence of erbium doped Zn₂SiO₄ glassceramics as phosphor in optoelectronic devices. *Journal of Alloys and Compounds*, 783, 441–447.
- El-Diasty, F., Abdel Wahab, F. A., & Abdel-Baki, M. (2006). Optical band gap studies on lithium aluminum silicate glasses doped with Cr³⁺ ions. *Journal of Applied Physics*, *100*(9), 3511.
- El-Mallawany, R. (2000). Structural interpretations on tellurite glasses. *Materials Chemistry and Physics*, *63*(2), 109–115.
- El-Moneim, A. A. (2014). Correlation between acoustical and structural parameters in some oxide glasses. *Journal of Non-Crystalline Solids*, *405*, 141–147.
- El-Moneim, A. A. (2019). Analysis and prediction of elastic moduli and Poisson's ratio in Li₂O-B₂O₃-V₂O₅ glasses under the substitution of V₂O₅ for B₂O₃.
 Physics and Chemistry of Glasses: European Journal of Glass Science and Technology Part B, 60(4), 146–156.
- El-rehim, A. F. A., Wahab, E. A. A., Halaka, M. M. A., & Shaaban, K. H. S. (2021). Optical properties of SiO₂-TiO₂-a₂O₃-Na₂O-Y₂O₃ glasses and a novel process of preparing the parent glass-ceramics. *Silicon*, *14*, 373–384.
- El Ghoul, J. (2018). Green and yellow luminescence properties of willemite Zn₂SiO₄ nanocomposites by sol–gel method. *Journal of Materials Science: Materials in Electronics*, *29*(4), 2999–3005.

- El Ghoul, J., Barthou, C., & El Mir, L. (2012a). Synthesis, structural and optical properties of nanocrystalline vanadium doped zinc oxide aerogel. *Physica E: Low-Dimensional Systems and Nanostructures*, *44*(9), 1910–1915.
- El Ghoul, J., Barthou, C., & El Mir, L. (2012b). Synthesis by sol-gel process, structural and optical properties of nanoparticles of zinc oxide doped vanadium. *Superlattices and Microstructures*, *51*(6), 942–951.
- El Ghoul, J., Omri, K., El Mir, L., Barthou, C., & Alaya, S. (2012c). Sol-gel synthesis and luminescent properties of SiO₂/Zn₂SiO₄ and SiO₂/Zn₂SiO₄:V composite materials. *Journal of Luminescence*, *132*(9), 2288–2292.
- El-Batal, H. A., Khalil, E. M. A., & Hamdy, Y. M. (2009). In vitro behavior of bioactive phosphate glass-ceramics from the system P₂O₅-Na₂O-CaO containing titania. *Ceramics International*, 35(3), 1195–1204.
- Elhadi, S. E., Liu, C., Zhao, Z., Li, K., & Zhao, X. (2018). Structure and optical properties of ZnO/Zn₂SiO₄ composite thin films containing Eu³⁺ ions. *Thin Solid Films*, 668, 1–8.
- Elkhoshkhany, N., Abbas, R., El-Mallawany, R., & Fraih, A. J. (2014). Optical Properties of quaternary TeO₂-ZnO-Nb₂O₅-Gd₂O₃ glasses. *Ceramics International*, 40(9 PART A), 14477–14481.
- Engku Ali, E. A. G., Matori, K. A., Saion, E., Aziz, S. H. A., Zaid, M. H. M., & Alibe, I. M. (2019). Effect of sintering temperatures on structural and optical properties of ZnO-Zn₂SiO₄ composite prepared by using amorphous SiO₂ nanoparticles. *Journal of the Australian Ceramic Society*, *55*(1), 115–122.
- Feldmann, C., Jüstel, T., Ronda, C. R., & Schmidt, P. J. (2003). Inorganic luminescent materials: 100 Years of research and application. *Advanced Functional Materials*, *13*(7), 511–516.
- Francis, A. A. (2004). Conversion of blast furnace slag into new glass-ceramic material. *Journal of the European Ceramic Society*, 24(9), 2819–2824.
- Gaafar, M. S., El-aal, N. S. A., Gerges, O. W., & El-amir, G. (2009). Elastic properties and structural studies on some zinc-borate glasses derived from ultrasonic, FT-IR and X-ray techniques. *Journal of Alloys and Compounds*, *475*, 535–542.
- Gao, X. D., Li, X. M., & Yu, W. D. (2004). Synthesis and optical properties of ZnO nanocluster porous films deposited by modified SILAR method. *Applied Surface Science*, *229*(1–4), 275–281.
- Gayathri Pavani, P., Sadhana, K., & Chandra Mouli, V. (2011). Optical, physical and structural studies of boro-zinc tellurite glasses. *Physica B: Condensed Matter*, *406*(6–7), 1242–1247.

- Ghobadi, N. (2013). Band gap determination using absorption spectrum fitting procedure. *International Nano Letters*, *3*(1), 2–5.
- Ha, M., Zaid, M., Amin, K., Hj, S., Aziz, A., Mohamed, H., Abdul, Z., Effendy, N., & Mustapha, I. (2016). Comprehensive study on compositional dependence of optical band gap in zinc soda lime silica glass system for optoelectronic applications. *Journal of Non-Crystalline Solids, 449*, 107– 112.
- Hafiz, M., Zaid, M., Matori, K. A., Hj, S., Aziz, A., & Zakaria, A. (2012). Effect of ZnO on the physical properties and optical band gap of soda lime silicate glass. *International Journal of Molecular Sciences*, 13, 7550–7558.
- Halimah, M. K., Daud, W. M., & Sidek, H. A. A. (2010). Elastic properties of TeO₂-B₂O₃-Ag₂O glasses. *Ionics, 16,* 807–813.
- Hamnabard, Z., Khalkhali, Z., Qazvini, S. S. A., Baghshahi, S., & Maghsoudipour, A. (2012). Preparation, heat treatment and photoluminescence properties of V-doped ZnO-SiO₂-B₂O₃ glasses. *Journal* of *Luminescence*, 132(5), 1126–1132.
- Hasnimulyati, L., Halimah, M. K., Zakaria, A., Halim, S. A., & Ishak, M. (2017). A Comparative study of the experimental and the theoretical elastic data of Tm³⁺ doped zinc borotellurite glass. *Materials Chemistry and Physics*. *192*(5), 228-234.
- Hassaan, M. Y., Osman, H. M., Hassan, H. H., Helal, M. A., & El-Deeb, A. S. (2019). Enhancing the electrical conduction in sodium borosilicate titanate glass doped with Nd or Gd ions to increase its optical absorption for smart windows applications. *Optik*, *185*, 477–485.
- He, Y., Chen, Q., Liu, H., Zhang, L., Wu, D., Lu, C., & Ouyang, W. (2019). Friction and Wear of MoO₃ / Graphene Oxide Modified Glass Fiber Reinforced Epoxy Nanocomposites. *Macromolecular Materials and Engineering*, 1900166, 1–11.
- Hou, L., Zuo, G., Shen, Y., Meng, Y., & Li, H. (2014). Effects of the replacing content of ZnBr₂ on the properties of ZnO-B₂O₃-SiO₂:Mn²⁺ glass-ceramics. *Ceramics International*, *40*(8), 13097–13103.
- Lozeman, JJA., Führer, P., Olthuis, W., & Odijk, M. (2020). Spectroelectrochemistry, the future of visualizing electrode processes by hyphenating electrochemistry with spectroscopic techniques. *Royal Society of Chemistry*. *145*, 2482-2509.
- Kamari, H. M., Aziz, S., & Talib, Z. A. (2005). Ultrasonic study and physical properties of borotellurite glasses. *American Journal of Applied Sciences 2, 11,* 1541-1546.

- Karmakar, B. (2016). Chapter 1 Fundamentals of glass and glass nanocomposites. In *Glass Nanocomposites: Synthesis, properties and applications* (pp. 3-53). William Andrew Applied Science Publishers.
- Kaur, A., Khanna, A., Pesquera, C., González, F., & Sathe, V. (2010). Preparation and characterization of lead and zinc tellurite glasses. *Journal* of Non-Crystalline Solids, 356(18–19), 864–872.
- Kaur, G., Kumar, M., Arora, A., Pandey, O. P., & Singh, K. (2011). Influence of Y₂O₃ on structural and optical properties of SiO₂-BaO-ZnO-xB₂O₃-(10x)Y₂O₃ glasses and glass ceramics. *Journal of Non-Crystalline Solids*, 357(3), 858–863.
- Kayahan, E. (2010). White light luminescence from annealed thin ZnO deposited porous silicon. *Journal of Luminescence*, *130*(7), 1295–1299.
- Kesavulu, C. R., & Jayasankar, C. K. (2012). Spectroscopic properties of Sm³⁺ ions in lead fluorophosphate glasses. *Journal of Luminescence*, *132*(10), 2802–2809.
- Khalkhali, Z., Hamnabard, Z., Qazvini, S. S. A., Baghshahi, S., & Maghsoudipour, A. (2012). Preparation, phase formation and photoluminescence properties of ZnO-SiO₂-B₂O₃ glasses with different ZnO/B₂O₃ ratios. *Optical Materials*, *34*(5), 850–855.
- Khan, H., Yerramilli, A. S., Oliveira, A. D., Alford, T. L., & Patience, B. G. S. (2019). Experimental methods in chemical engineering: X-ray diffraction spectroscopy XRD. *The Canadian Journal of Chemical Engineering*, *98*(6), 1255-1266.
- Kohara, S., Suzuya, K., Takeuchi, K., Loong, C.-K., Grimsditch, M., Weber, J. K. R., Tangeman, J. A., & Key, T. S. (2004). Glass formation at the limit of insufficient network formers. *Science*, 303, 1649–1652.
- Krsmanovi, R. M., Anti, Ž., & Drami, M. D. (2011). Structural, spectroscopic and crystal field analyses of Ni²⁺ and Co²⁺ doped Zn₂SiO₄ powders. *Applied Physics A*, *104*, 483–492.
- Kuang, J., Liu, Y., & Lei, B. (2006). Effect of RE³⁺ as a co-dopant in long-lasting phosphorescence CdSiO₃:Mn²⁺ (RE=Y, La, Gd, Lu). *Journal of Luminescence*, *118*(1), 33–38.
- Kullberg, A. T. G., Lopes, A. A. S., Veiga, J. P. B., Lima, M. M. R. A., & Monteiro, R. C. C. (2016). Formation and crystallization of zinc borosilicate glasses: Influence of the ZnO/B₂O₃ ratio. *Journal of Non-Crystalline Solids*, 441, 79– 85.
- Kurama, S., Kara, A., & Kurama, H. (2006). The effect of boron waste in phase and microstructural development of a terracotta body during firing. *Journal* of the European Ceramic Society, 26(4–5), 755–760.

- Laila, S., Suraya, A. K., & Yahya, A. K. (2014). Effect of glass network modification on elastic and structural properties of mixed electronic-ionic 35V₂O₅-(65-x)TeO₂-(x)Li₂O glass system. *Chalcogenide Letters*, *11*(2), 91– 104.
- Lee, C. S., Amin Matori, K., Ab Aziz, S. H., Kamari, H. M., Ismail, I., & Mohd Zaid, M. H. (2017). Comprehensive study on elastic moduli prediction and correlation of glass and glass ceramic derived from waste rice husk. *Advances in Materials Science and Engineering*, 2017, 1-10.
- Leyva-Porras, C., Cruz-Alcantar, P., Espinosa-Sol, V., & Saavedra-Leos, M. Z. (2019). Application of differential scanning calorimetry (DSC) and modulated differential scanning calorimetry (MDSC) in food and drug industries. *Polymers*, *12*(5), 1–21.
- Li, H. C., Wang, D. G., Hu, J. H., & Chen, C. Z. (2013). Effect of the partial substitution of K₂O, MgO, B₂O₃ for CaO on crystallization, structure and properties of Na₂O-CaO-SiO₂-P₂O₅ system glass-ceramics. *Materials Letters*, *106*, 373–376.
- Li, Z. Y., & Zhang, Z. Q. (2000). Fragility of photonic band gaps in inverse-opal photonic crystals. *Physical Review B Condensed Matter and Materials Physics*, 62(3), 1516–1519.
- Lin, J., Sänger, D. U., Mennig, M., & Bärner, K. (2000). Sol-gel deposition and characterization of Mn²⁺-doped silicate phosphor films. *Thin Solid Films*, *360*(1–2), 39–45.
- Liu, X., Zhou, J., Zhou, S., Yue, Y., & Qiu, J. (2018). Transparent glass-ceramics functionalized by dispersed crystals. *Progress in Materials Science*, 97, 38– 96.
- Loiko, P., Dymshits, O., Volokitina, A., Alekseeva, I., Shemchuk, D., Tsenter, M., Bachina, A., Khubetsov, A., Vilejshikova, E., Petrov, P., Baranov, A., & Zhilin, A. (2018). Structural transformations and optical properties of glassceramics based on ZnO, β - and α -Zn₂SiO₄ nanocrystals and doped with Er₂O₃ and Yb₂O₃: Part I. The role of heat-treatment. *Journal of Luminescence*, *202*, 47–56.
- Lu, Q., & Yun, G. (2013). Facile one-step solid-state reaction route to synthesize ordered mesoporous β-Zn₂SiO₄-SiO₂ nanocomposites. *Ceramics International*, *39*(4), 3533–3538.
- Luo, Y., Zhang, S., Nikl, M., Ye, S., & Hu, R. (2016). Luminescence of rare-earth ions and intrinsic defects in Gd₂O₃ matrix. *Journal of Physics: Conference Series*, *741*, 012089.
- Lushchik, A., Kirm, M., Lushchik, C., Martinson, I., & Zimmerer, G. (2000). Luminescence of free and self-trapped excitons in wide-gap oxides. *Journal* of Luminescence, 89, 232–234.

- Ma, J. G., Liu, Y. C., Xu, C. S., Liu, Y. X., Shao, C. L., Xu, H. Y., Zhang, J. Y., Lu, Y. M., Shen, D. Z., & Fan, X. W. (2005). Preparation and characterization of ZnO particles embedded in SiO₂ matrix by reactive magnetron sputtering. *Journal of Applied Physics*, *97*(10), 3509.
- Maheshwary, Singh, B. P., & Singh, R. A. (2015). Color tuning in thermally stable Sm³⁺-activated CaWO₄ nanophosphors. *New Journal of Chemistry*, *39*(6), 4494–4507.
- Mallur, S. B., Czarnecki, T., Adhikari, A., & Babu, P. K. (2015). Compositional dependence of optical band gap and refractive index in lead and bismuth borate glasses. *Materials Research Bulletin*, 68, 27–34.
- Martino, D. D., Chiodini, N., Fasoli, M., Moretti, F., Vedda, A., Baraldi, A., Buffagni, E., Capelleti, R., Mazzera, M., Nikl, M., Angella, G., & Azzoni, C. B. (2008). Gd-incorporation and luminescence properties in sol-gel silica glasses. *Journal of Non-Crystalline Solids*, 354(32), 3817–3823.
- Marzouk, S. Y., & Gaafar, M. S. (2007). Ultrasonic study on some borosilicate glasses doped with different transition metal oxides. *Solid State Communications*, *144*(10–11), 478–483.
- Matori, K. A., Hafiz, M., Zaid, M., Quah, H. J., Hj, S., Aziz, A., Wahab, Z. A., & Ghazali, M. (2015). Studying the effect of ZnO on physical and elastic properties of (ZnO)x P₂O₅)1-x glasses using nondestructive ultrasonic method. *Advances in Materials Science and Engineering*, 2015, 1-6.
- Menon, R., Gupta, V., Tan, H. H., Sreenivas, K., & Jagadish, C. (2011). Origin of stress in radio frequency magnetron sputtered zinc oxide thin films. *Journal of Applied Physics*, *109*(6), 4905.
- Min, B., Koo, H., & Seok, B. (2006). Generation of glass SiO₂ structures by various cooling rates : A molecular-dynamics study. *37*, 203–208.
- Monteiro, R. C. C., Lopes, A. A. S., Lima, M. M. R. A., & Veiga, J. P. B. (2018). Thermal characteristics and crystallization behavior of zinc borosilicate glasses containing Nb₂O₅. *491*, 124–132.
- Munjanja, B., & Sanganyado, E. (2015). UV-Visible absorption, fluorescence ,and chemiluminescence spectroscopy. *CRC Press*, (December 2016), 572–583.
- Mustapha, S., Ndamitso, M. M., Abdulkareem, A. S., Tijani, J. O., Shuaib, D. T., Mohammed, A. K., & Sumaila, A. (2019). Comparative study of crystallite size using Williamson-Hall and Debye-Scherrer plots for ZnO nanoparticles. Advances in Natural Sciences: Nanoscience and Nanotechnology, 10(4).

Nemova, G., & Kashyap, R. (2012). Laser cooling with Tm³⁺-doped oxy-fluoride glass ceramic. *Journal of the Optical Society of America B*, 29(11), 3034.
 Nien, Y. T., Chen, Y. L., Chen, I. G., Hwang, C. S., Su, Y. K., Chang, S. J., &

Juang, F. S. (2005). Synthesis of nano-scaled yttrium aluminum garnet phosphor by co-precipitation method with HMDS treatment. *Materials Chemistry and Physics*, *93*(1), 79–83.

- Nuraidayani, E., Ab, H., Sidek, A., Kamari, H. M., Ha, M., Zaid, M., ... Elmallawany, R. (2019). Optik synthesis and green luminescence of low cost Er₂O₃ doped zinc silicate glass-ceramics as laser materials. *184*(April), 480–484.
- Omar, N. A. S., Fen, Y. W., & Matori, K. A. (2017). Europium doped low cost Zn₂SiO₄ based glass ceramics: A study on fabrication, structural, energy band gap and luminescence properties. *Materials Science in Semiconductor Processing*, *61*, 27–34.
- Omri, K., El Ghoul, J., Alyamani, A., Barthou, C., & El Mir, L. (2013). Luminescence properties of green emission of SiO₂/Zn₂SiO₄:Mn nanocomposite prepared by sol-gel method. *Physica E: Low-Dimensional Systems and Nanostructures*, *53*, 48–54.
- Omri, K., Lemine, O. M., & El Mir, L. (2017). Mn doped zinc silicate nanophosphor with bifunctionality of green-yellow emission and magnetic properties. *Ceramics International*, *43*(8), 6585–6591.
- Özgür, Ü., Alivov, Y. I., Liu, C., Teke, A., Reshchikov, M. A., Doğan, S., Avrutin, V., Cho, S. J., & Morko, H. (2005). A comprehensive review of ZnO materials and devices. *Journal of Applied Physics*, *98*(4), 1–103.
- Paulen, R. C., McClenaghan, M. B., & Hicken, A. K. (2013). Regional and local ice-flow history in the vicinity of the Izok Lake Zn-Cu-Pb-Ag deposit, Nunavut. Canadian Journal of Earth Sciences, 50(12), 1209–1222.
- Qazvini, S. S. A., Hamnabard, Z., Khalkhali, Z., Baghshahi, S., & Maghsoudipour, A. (2012). Photoluminescence and microstructural properties of SiO₂-ZnO-B₂O₃ system containing TiO₂ and V₂O₅. *Ceramics International*, *38*(2), 1663–1670.
- Rajendran, V., Palanivelu, N., Modak, D. K., & Chaudhuri, B. K. (2000). Ultrasonic investigation on ferroelectric BaTiO₃ doped 80V₂O₅-20PbO oxide glasses. *Physica Status Solidi (a), 180*(2), 467–477.
- Rao, V. V. S. J., Kannan, E., Prakash, R. V, Balasubramaniam, K. (2013). Fatigue damage characterization using surface acoustic wave nonlinearity in aluminum alloy AA7175-T7351. *Journal of Applied Physics*, 104, 123508(2008).
- Raunak, K. T., Bisen, D. P., Upadhyay, K., Sahu, M., Sahu, I. P., & Brahmen, N. (2015). Comparison of emitted color by pure Gd₂O₃ prepared by two different methods by CIE coordinates. *Superlattices and Microstructures*, *88*, 382–388.

- Rawlings, R. D., Wu, J. P., & Boccaccini, A. R. (2006). Glass-ceramics: Their production from wastes-A Review. *Journal of Materials Science*, 41(3), 733–761.
- Ruan, H. B., Fang, L., Li, D. C., Saleem, M., Qin, G. P., & Kong, C. Y. (2011). Effect of dopant concentration on the structural, electrical and optical properties of Mn-doped ZnO films. *Thin Solid Films*, *519*(15), 5078–5081.
- Saddeek, Y. B., & Gaafar, M. S. (2009). Physical and structural properties of some bismuth borate glasses. *Materials Chemistry and Physics*, 115(1), 280–286.
- Saddeek, Y. B., & Latif, L. A. El. (2004). Effect of TeO2 on the elastic moduli of sodium borate glasses. *Physica B: Condensed Matter*, 348(1–4), 475–484.
- Sagadevan, S., & Murugasen, P. (2014). Studies on optical , mechanical and electrical properties of organic nonlinear optical p-toluidine ptoluenesulfonate single. *Journal of Crystallization Process and Technology*, (April), 99–110.
- Sahu, I. P., Bisen, D. P., Brahme, N., & Tamrakar, R. K. (2016). Studies on the luminescence behavior of SrCaMgSi₂O₇:Eu³⁺ phosphor by solid state reaction method. *Journal of Materials Science: Materials in Electronics*, 27(2), 1828–1839.
- Salinigopal, M. S., Gopakumar, N., & Anjana, P. S. (2019). Alkaline earth based borosilicate glasses as sealants in solid oxide fuel cell applications. *Silicon*, *12*, 101-107.
- Salinigopal, M. S., Gopakumar, N., Anjana, P. S., & Pandey, O. P. (2020). Synthesis and characterization of 50BaO-(5-x)Al₂O₃-xR₂O₃-30B₂O₃-15SiO₂(R = Nd, Gd) glass-ceramics. *Journal of Non-Crystalline Solids, 535*, 119956.
- Samsudin, N. F., Matori, K. A., Liew, J. Y. C., Fen, Y. W., Mohd Zaid, M. H., & Alassan, Z. N. (2015). Investigation on structural and optical properties of willemite doped Mn²⁺ based glass-ceramics prepared by conventional solid-state method. *Journal of Spectroscopy*, 2015, 1-7.
- Sasmal, N., Garai, M., & Karmakar, B. (2016). Influence of Ce, Nd, Sm and Gd oxides on the properties of alkaline-earth borosilicate glass sealant. *Journal of Asian Ceramic Societies*, *4*(1), 29–38.
- Sayyed, M. I., Askin, A., Zaid, M. H. M., Olukotun, S. F., Uddin, M., Tishkevich, D. I., & Bradley, D. A. (2021). Radiation shielding and mechanical properties of Bi₂O₃ – Na₂O – TiO₂ – ZnO – TeO₂ glass system. *Radiation Physics and Chemistry, 186*, 109556.
- Sayyed, M. I., Kumar, A., Tekin, H. O., Kaur, R., Singh, M., & Agar, O. (2020). Evaluation of gamma-ray and neutron shielding features of heavy metals doped Bi₂O₃-BaO-Na₂O-MgO-B₂O₃ glass systems. *Progress in Nuclear Energy*, *118*, 103118.

- Schabbach, L. M. C., Andreola, F., Karamanova, E., Lancellotti, I., Karamanov, A., & Barbieri, L. (2011). Integrated approach to establish the sintercrystallization ability of glasses from secondary raw material. *Journal of Non-Crystalline Solids*, 357(1), 10–17.
- Semnani, D. (2017). Geometrical characterization of electrospun nanofibers. In *electrospun nanofibers*. (pp. 151-180), Woodhead Publishing Series in Textiles.
- Sessolo, M., & Bolink, H. J. (2011). Hybrid organic-inorganic light-emitting diodes. *Advanced Materials*, 23(16), 1829–1845.
- Sharma, P., & Bhatti, H. S. (2011). Effects of dopant concentrations and firing temperatures on decay kinetics of manganese doped willemite nanopowders. *Physica B: Physics of Condensed Matter, 406*(22), 4188–4194.
- Shen, Y., Hou, L., Zuo, G., Li, F., & Meng, Y. (2014). Preparation of ZnO–B₂O₃– SiO₂:Mn²⁺ optical-storage glass–ceramics with different ZnF₂ dopant by sol–gel method. *Journal of Sol-Gel Science and Technology*, 73(1), 192– 198.
- Shenzhen, T. (2015). Emission spectra of Tb³⁺:Zn₂SiO₄ and Eu³⁺:Zn₂SiO₄ solgel powder phosphors. *Journal of Spectroscopy and Dynamics*, *4*, 1–8.
- Sidek, H. A. A., El-Mallawany, R., Hariharan, K., & Rosmawati, S. (2014). Effect of concurrent ZnO addition and AIF₃ reduction on the elastic properties of tellurite based glass system. *Advances in Condensed Matter Physics*, 2014, 1-7.
- Sidek, H. A. A., El-Mallawany, R., Matori, K. A., & Halimah, M. K. (2016). Effect of PbO on the elastic behavior of ZnO-P₂O₅ glass systems. *Results in Physics*, *6*, 449–455.
- Sidek, H. A. A., Rosmawati, S., Azmi, B. Z., & Shaari, A. H. (2013). Effect of ZnO on the thermal properties of tellurite glass. *Advances in Condensed Matter Physics*, 2013, 1-6.
- Singla, M. L., Shafeeq M, M., & Kumar, M. (2009). Optical characterization of ZnO nanoparticles capped with various surfactants. *Journal of Luminescence*, *129*(5), 434–438.
- Sohn, K. S., Cho, B., & Park, H. D. (1999). Excitation energy-dependent photoluminescence behavior in Zn₂SiO₄:Mn phosphors. *Materials Letters*, *41*(6), 303–308.
- Stefan, R., Culea, E., & Pascuta, P. (2012). The effect of copper ions addition on structural and optical properties of zinc borate glasses. *Journal of Non-Crystalline Solids*, 358(4), 839–846.

- Su, K., Tilley, T. D., & Sailor, M. J. (1996). Molecular and polymer precursor routes to manganese-doped zinc orthosilicate phosphors. *Journal of Amaerican Chemical Society*, 118(614, 3459–3468.
- Suzuki, T., & Ohishi, Y. (2004). Broadband 1400 nm emission from Ni²⁺ in zinc alumino silicate glass. *Applied Physics Letters*, *84*(19), 3804–3806.
- Tagiara, N. S., Palles, D., Simandiras, E. D., Psycharis, V., Kyritsis, A., & Kamitsos, E. I. (2017). Synthesis, thermal and structural properties of pure TeO₂ glass and zinc-tellurite glasses. *Journal of Non-Crystalline Solids*, 457, 116–125.
- Takesue, M., Hayashi, H., & Smith, R. L. (2009). Thermal and chemical methods for producing zinc silicate (willemite): A review. *Progress in Crystal Growth and Characterization of Materials*, *55*(3–4), 98–124.
- Tanabe, S., Hayashi, H., Hanada, T., & Onodera, N. (2002). Fluorescence properties of Er³⁺ ions in glass ceramics containing LaF₃ nanocrystals. *Optical Materials*, *19*(3), 343–349.
- Tarafder, A., Molla, A. R., Dey, C., & Karmakar, B. (2013). Thermal, structural, and enhanced photoluminescence properties of Eu³⁺-doped transparent willemite glass-ceramic nanocomposites. *Journal of the American Ceramic Society*, *96*(8), 2424–2431.
- Tarafder, A., Molla, A. R., Mukhopadhyay, S., & Karmakar, B. (2012). Synthesis and properties of SrBi₂ Ta₂O₉-based glass-ceramics modified with Eu³⁺. *Journal of the American Ceramic Society*, *95*(6), 1851–1857.
- Tarafder, A., Molla, A. R., Mukhopadhyay, S., & Karmakar, B. (2014). Fabrication and enhanced photoluminescence properties of Sm³⁺-doped ZnO-Al₂O₃-B₂O₃-SiO₂ glass derived willemite glass-ceramic nanocomposites. *Optical Materials*, *36*(9), 1463–1470.
- Tauc, J. (1968). Optical properties and electronic structure of amorphous Ge and Si. *Materials Research Bulletin*, *3*, 37–46.
- Tick, P. A., Borrelli, N. F., Cornelius, L. K., & Newhouse, M. A. (1995). Transparent glass ceramics for 1300 nm amplifier applications. *Journal of Applied Physics*, *78*(11), 6367–6374.
- Va, P., Nekvindová, P., Michalcová, A., & Malinský, P. (2019). The influence of copper and silver in various oxidation states on the photoluminescence of Ho³⁺/Yb³⁺ doped zinc-silicate glasses. *Optical Materials*, *91*, 253–260.
- Van Der Kolk, E., Dorenbos, P., Van Eijk, C. W. E., Bechtel, H., Jüstel, T., Nikol, H., Ronda, C. R., & Wiechert, D. U. (2000). Optimised co-activated willemite phosphors for application in plasma display panels. *Journal of Luminescence*, 87, 1246–1249.

- Van, P. Š., & Klement, R. (2016). Photoluminescence of (ZnO)x-z(SiO₂)y:(MnO)z green phosphors prepared by direct thermal synthesis : The effect of ZnO/SiO₂ ratio and Mn²⁺ concentration on luminescence. *Ceramics International, 42(15),* 16852-16860.
- Wahab, S. A. A., Matori, K. A., Aziz, S. H. A., Zaid, M. H. M., Kechik, M. M. A., Azman, A. Z. K., Khaidir, R. E. M., Khiri, M. Z. A., & Effendy, N. (2020). Effect of ZnO on the phase transformation and optical properties of silicate glass frits using rice husk ash as a SiO₂ source. *Journal of Materials Research and Technology*, 9(5), 11013–11021.
- Wan, J., Wang, Z., Chen, X., Mu, L., Yu, W., & Qian, Y. (2006). Controlled synthesis and relationship between luminescent properties and shape/crystal structure of Zn₂SiO₄:Mn²⁺ phosphor. *Journal of Luminescence*, *121*(1), 32–38.
- Watanabe, M. Ã., & Hayashi, T. (2005). Time-resolved study of self-trapped exciton luminescence in anatase TiO₂ under two-photon excitation. *Journal of Luminescence*, *112*, 88–91.
- Wegh, R. T., Donker, H., & Meijerink, A. (1997). Vacuum-ultraviolet spectroscopy and quantum cutting for Gd³⁺ in LiYF₄. *Physical Review B*, *56*(21), 841– 848.
- Yan, B., & Huang, H. (2007). Matrix-inducing synthesis and luminescence of Zn₂SiO₄:xTb³⁺ submicrometer phosphors derived from the sol-gel assembling of different multicomponent hybrid precursors. *Journal of Alloys and Compounds*, *429*(1–2), 338–342.
- Yao, Y., & Zhou, Z. (2016). Photoluminescence characteristics of a novel red phosphor Ba₂Si₄O10:Eu³⁺: structural effect and concentration quenching mechanism. *Journal of Luminescence*, *179*, 408–412.
- Yaowakulpattana, P., Kondo, S., Kadono, K., & Wakasugi, T. (2015). Effect of B₂O₃ on crystallization behavior of ZnO-Al₂O₃-SiO₂ glasses. *Journal of the Ceramic Society of Japan*, 123(1434), 96–99.
- Yekta, B. E., Alizadeh, P., & Rezazadeh, L. (2007). Synthesis of glass-ceramic glazes in the ZnO-Al₂O₃-SiO₂-ZrO₂ system. *Journal of the European Ceramic Society*, *27*(5), 2311–2315.

Yoshiyuki, K., Adachi, H., & Minami, T. (1995). Electronic states of transition metal ions in silicate glasses. *Journal of Non-Crystalline Solids, 192-193*, 316–320.

Yu, Y., Yao, Y., Yan, H., Wang, R., Zhang, Z., Sun, X., Zhao, L., Ao, X., Xie, Z., & Wu, Q. (2016). A tumor-specific MicroRNA recognition system facilitates the accurate targeting to tumor cells by magnetic nanoparticles. *Molecular Therapy - Nucleic Acids*, *5*, e318.

- Zachariasen, W. H. (1932). The atomic arrangement in glass. *Journal of the American Chemical Society*, *54*(10), 3841–3851.
- Zaid, M. H. M., Matori, K. A., Abdul Aziz, S. H., Kamari, H. M., Wahab, Z. A., Fen, Y. W., & Alibe, I. M. (2016). Synthesis and characterization of low cost willemite based glass–ceramic for opto-electronic applications. *Journal of Materials Science: Materials in Electronics*, 27(11), 11158–11167.
- Zamratul, M. I. M., Zaidan, A. W., Khamirul, A. M., Nurzilla, M., & Halim, S. A. (2016). Formation, structural and optical characterization of neodymium doped-zinc soda lime silica based glass. *Results in Physics*, *6*, 295–298.
- Zeng, H., Duan, G., Li, Y., Yang, S., Xu, X., & Cai, W. (2010). Blue luminescence of ZnO nanoparticles based on non-equilibrium processes: Defect origins and emission controls. *Advanced Functional Materials*, 20(4), 561–572.
- Zeng, Y., Qin, X., Jiang, S., Zhang, G., & Zhang, L. (2011). Effect of BaF₂ addition on crystallization kinetics and dielectric properties of B₂O₃-Nb₂O₅-SrO-BaO glass-ceramics. *Journal of the American Ceramic Society*, *94*(2), 469–473.
- Zhang, Y., Han, K., Cheng, T., & Fang, Z. (2007). Synthesis , characterization , and photoluminescence property of LaCO₃OH microspheres. *Inorganic Chemistry*, *46*(11), 827–831.
- Zhou, S., Li, C., Yang, G., Bi, G., Xu, B., Hong, Z., Miura, K., Hirao, K., & Qiu, J. (2013). Self-limited nanocrystallization-mediated activation of semiconductor nanocrystal in an amorphous solid. Advanced Functional Materials, 23(43), 5436–5443.